

Thermochemistry Review

Multiple Choice

Identify the choice that best completes the statement or answers the question.

Use the following information to answer the next _ questions.

Decaying matter may provide energy directly through combustion or indirectly through decomposition to form methane, which can then be burned.



- _____ 1. The burning of methane is an example of an
- exothermic reaction in which carbon is oxidized.
 - exothermic reaction in which carbon is reduced.
 - endothermic reaction in which carbon is oxidized.
 - endothermic reaction in which carbon is reduced.
- _____ 2. The molar enthalpy of combustion for methane is
- 560.5 kJ/mol
 - 710.1 kJ/mol
 - 802.3 kJ/mol
 - 951.9 kJ/mol
- _____ 3. An equation that represents an endothermic change is
- $\text{C}_2\text{H}_5\text{OH}(\text{l}) + 2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$
 - $2 \text{C}(\text{s}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g})$
 - $\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
 - $2 \text{C}(\text{s}) + 3 \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$
- _____ 4. Sulfuric acid is used to dissolve copper ore in many mining operations. Sulfuric acid may be produced on site by burning sulfur and dissolving the products in water. The overall equation for this reaction is
- $$\text{S}_8(\text{g}) + 12 \text{O}_2(\text{g}) + 8 \text{H}_2\text{O}(\text{l}) \rightarrow 8 \text{H}_2\text{SO}_4(\text{aq}) + 4225.6 \text{ kJ}$$
- In this process, what is the enthalpy of reaction for the production of one mole of sulfuric acid?
- +4225.6 kJ
 - +528.2 kJ
 - 528.2 kJ
 - 4225.6 kJ
- _____ 5. Which of the following statements is **false**?
- Light energy is converted into chemical energy during photosynthesis.
 - Chemical energy is converted into heat and light energy during the combustion of hydrocarbons.
 - Light energy is converted into chemical energy to produce a firefly's glow.
 - Chemical energy is converted into electrical energy as a voltaic cell discharges.

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The Princeton Review



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in the field of Thermochemistry of Alloys convened for a NATO Advanced Study Institute at Kiel in August 1987 to discuss these developments This book collects most of the lectures and seminar papers given at the Institute *General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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CALCULATIONS K. V. NARAYANAN,B. LAKSHMIKUTTY,2006-01-01 This textbook is designed for undergraduate courses in chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering safety engineering and industrial chemistry The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE **AP**

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