

6.023×10^{23}
particles

in terms of
particles

in terms of
volume

22.4 L of a
gas at N.T.P

1 Mole

1 mole of oxygen atoms
= 6.023×10^{23} atoms

1 mole of oxygen molecules
= 6.023×10^{23} molecules

in terms of
mass

1 gram atom of
elements

12g
C

23g
Na

1 gram molecule
of substance

18g
 H_2O

44g
 CO_2

1 gram formula
weight of
substance

58.5 g
NaCl

100g
 CaCO_3

Unit 7 Mole Relationships

Judd E. Hollander

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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