

Unit 10 Stoichiometry Test 1: Short Answers

Answer all questions including units and balanced equations where appropriate.

1. How many moles of aluminium are produced when completely with 2.5 mol of FeCl_3 ?
$$2\text{Al} + 3\text{FeCl}_2 \rightarrow 3\text{Fe} + \text{Al}_2\text{Cl}_6$$
2. How many grams of FeCl_3 are produced when 2.5 moles of aluminium with oxygen to form Fe_2O_3 ?
$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$
3. The reaction below shows that decomposition of FeCl_3 to form iron and chlorine gas.
$$2\text{FeCl}_3 \rightarrow 2\text{Fe} + 3\text{Cl}_2$$

How many grams of chlorine are produced when 25.0g of FeCl_3 is decomposed?
4. An element that has been calculated to produce 2.0g of AgCl actually produces 2.2g of AgCl . What is the percent yield?
5. Balance the equation and determine the limiting reagent when 25.0g of nitrogen react with 25.0g of hydrogen according to this equation.
$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$
6. Limiting Reagent is
7. Amount of NH_3 produced is
8. If you actually obtain 2.2g of AgCl , what is the percent yield?

Unit 8 Test Stoichiometry

AW Chickering



Unit 8 Test Stoichiometry:

Reviewing **Unit 8 Test Stoichiometry**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Unit 8 Test Stoichiometry**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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