

Chemical Equilibrium Assignment Quiz

A.	Briefly describe or define and give an example of (a) a weak electrolyte. (b) a Brønsted-Lowry acid. (c) the conjugate acid of a Brønsted-Lowry base. (d) neutralization, in terms of the Brønsted-Lowry concept. (e) an amphiprotic solvent. (f) a zwitterion. (g) autoprotolysis. (h) a strong acid. (i) Le Châtelier's principle. (j) the common-ion effect.
B.	Identify the acid on the left and its conjugate base on the right in the following equations: (a) $\text{HOCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OCl}^-$ (b) $\text{HONH}_2 + \text{H}_2\text{O} \rightleftharpoons \text{HONH}_3^+ + \text{OH}^-$ (c) $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{NH}_3 + \text{H}_3\text{O}^+$ (d) $2\text{HCO}_3^- \rightleftharpoons \text{H}_2\text{CO}_3 + \text{CO}_3^{2-}$ (e) $\text{PO}_4^{3-} + \text{H}_2\text{PO}_4^- \rightleftharpoons 2\text{HPO}_4^{2-}$ Identify the base on the left and its conjugate acid on the right in the equations for a-e above
C.	Write the equilibrium-constant expressions and obtain numerical values for each constant in (a) the basic dissociation of aniline, $\text{C}_6\text{H}_5\text{NH}_2$. (b) the acidic dissociation of hypochlorous acid, HClO. (c) the acidic dissociation of methyl ammonium hydrochloride, $\text{CH}_3\text{NH}_3\text{Cl}$.
D.	Calculate the solubility-product constant for each of the following substances, given that the molar concentrations of their saturated solutions are as indicated: (a) AgSeCN ($2.0 \times 10^{-8} \text{ M}$; products are Ag^+ and SeCN^-). (b) RaSO_4 ($6.6 \times 10^{-6} \text{ M}$).
E.	Use Equation 10-5 to calculate the activity coefficient of (a) Fe^{3+} at $\mu = 0.062$. (b) Pb^{2+} at $\mu = 0.042$. (c) Ce^{4+} at $\mu = 0.070$. (d) Sn^{4+} at $\mu = 0.045$.
F.	For a solution in which $\mu = 8.0 \times 10^{-2}$, calculate K_{sp}^* for (a) AgSCN. (b) PbI_2.

Unit Chemical Equilibrium Assignment Answers

Albert A Gayle



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