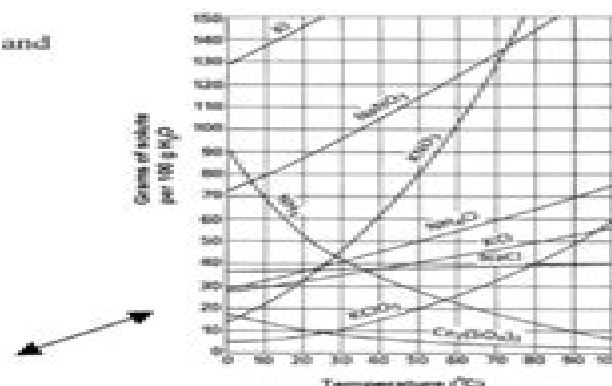


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- Review learning targets for Unit 12
- Solute, solvent, soln
- Like dissolves like
- Colligative properties (freezing point depression, boiling point elevation, Van't Hoff Factor)
- Saturated, unsaturated, supersaturated
- 4 types of mixtures and examples and examples of each
- What are the differences between the solubilities of gases and solids
- Naming and formula writing
- Steps to make a solution
- Steps to dilute a solution

Solution	Supersaturated, Saturated or Unsaturated?	If unsaturated: How much more solute can dissolve in the solution?
a solution that contains 95g of NaNO_3 at 30°C (in 100 mL H_2O)		
a solution that contains 50g of NH_4Cl at 50°C (in 100 mL H_2O)		
a solution that contains 20g of KClO_3 at 50°C (in 100 mL H_2O)		
a solution that contains 70g of KI at 0°C (in 100 mL H_2O)		



- Calculate the **molarity (M)** of a solution prepared by dissolving 175 g of KNO_3 in 750 mL of water.
- Calculate the **molarity (M)** of a solution prepared by dissolving 235.0 g of CaCl_2 in 650.0 mL of water.
- How many mL of 3.0M HCl are needed to make 300.0 mL of 0.10M HCl ?

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