Question 1: Water at 20° C falls from a height of 854 m. If the whole energy is used in increasing the temperature, find out the final temperature. Specific heat of water is 4200 J/K kg.

Given that

$$T_1 = 20^{\circ}C = 20 + 273 = 293 \,\mathrm{K}$$

height = $h = 854 \,\mathrm{m}$

Specific heat of water = 4200 JK 'kg'

Final temperature = $T_2 = ?$

P.E = mgh — 0

alution:

we know that

Solution:

DE KNOW THAT

$$\Delta Q = C M \Delta T - Q$$

Whole energy is used in increasing the temperature (Given)

From
$$\bigcirc$$
 and \bigcirc
 C $\nearrow AT = \gamma / g h$
 $C \wedge T = g h$
 $C \wedge T_2 - T_1 \rangle - g h$

$$T_{2} - T_{1} = \frac{9h}{2}$$

$$T_{2} = \frac{9h}{2} + T_{1}$$

$$= 9.8 \times 854$$

$$= \frac{9.8 \times 854}{4200} + 293$$

$$= \frac{8369.2}{4200} + 293 = 1.99 + 293$$

$$T_2 = 294.99 \text{ K} = 295 - 273 °C$$

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