

Unit 2 Worksheet: St. Michael's Concentration

Part A: Dissolving Metal Concentration

1. Cadmium chloride has a solubility in water for pressures at 1.000 mol/L. Explain what this concentration tells you about the CdCl_2 solution. What might a chemical use this method?

2. A 20.00 g sample of potassium nitrate is dissolved in enough water to make 200.0 mL of solution. Calculate the molar concentration of the potassium nitrate.

3. Find the molarity of a solution made from 2.50 g of CaCl_2 dissolved in enough water to make 0.100 L.

4. An unknown metal reacts with hydrochloric acid and produces hydrogen gas. Dissolving 0.20 g of metal produces 0.040 g of hydrogen gas. Calculate the molarity of the metal in the solution.

5. Calculate the molar concentration of a solution of sodium at 0.20 M NaCl solution from 0.2 g of NaCl (2.0)?

Unit 9 Worksheet 3 Molar Concentration Stoichiometry

Jicheng Xie



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