

Unit 10 Stoichiometry Test 1: Short Answers

Answer all questions including units and balanced equations where appropriate.

1. How many moles of aluminium are produced when completely with 2.5 mol of FeCl_3 ?

$$2\text{Al} + 3\text{FeCl}_2 \rightarrow 3\text{Fe} + 2\text{AlCl}_3$$

2. How many grams of FeCl_3 are produced when 2.5 moles of aluminium reacts completely with FeCl_2 ?

$$2\text{Al} + 3\text{FeCl}_2 \rightarrow 3\text{Fe} + 2\text{AlCl}_3$$

3. The reaction below shows that decomposition of FeCl_3 is complete. How many grams of oxygen are produced when 20.7g FeCl_3 is heated?

$$2\text{FeCl}_3 \rightarrow 2\text{Fe} + 3\text{Cl}_2$$

4. A reaction that has been calculated to produce 2.0g of AgCl actually produces 2.2g of AgCl . What is the percent yield?

5. Balance the equation and determine the limiting reagent when 20.0g of nitrogen reacts with 20.0g of hydrogen according to this equation.

$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$

6. Limiting Reagent is
7. Amount of NH_3 produced is
8. How actually when 20.0g of N_2 reacts with 20.0g of H_2 ?

Unit 10 Stoichiometry Test Review

**M. Michael Miller, Deborah A.
Kramer, G. Oliver Vagt**



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