

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this same solution had been added to the NaCl solution. This would be identical to the first solution.
 Their respective volumes (1.5 moles NaCl) for every 1L of solution, it would be the same for the second solution, which would produce a 1.5M solution.
2. A 10.0 g sample of potassium chloride is dissolved in enough water to make 100 mL of solution. Chlorine has a molar mass of 35.45 g/mol and potassium has a molar mass of 39.10 g/mol.

$$10.0 \text{ g} \times \frac{1 \text{ mol}}{74.55 \text{ g}} = 0.134 \text{ mol}$$

$$0.134 \text{ mol} \times \frac{1 \text{ L}}{0.100 \text{ L}} = 1.34 \text{ M}$$

3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.5 L.

$$25.0 \text{ g} \times \frac{1 \text{ mol}}{110.98 \text{ g}} = 0.225 \text{ mol}$$

$$0.225 \text{ mol} \times \frac{1 \text{ L}}{0.500 \text{ L}} = 0.450 \text{ M}$$

4. An unknown ionic solution contains "chlorine" and sodium in a 1:1 ratio. A 10.0 g sample of this solution is dissolved in enough water to make 0.5 L. Find the molarity of the solution.

$$10.0 \text{ g} \times \frac{1 \text{ mol}}{74.55 \text{ g}} = 0.134 \text{ mol}$$

$$0.134 \text{ mol} \times \frac{1 \text{ L}}{0.500 \text{ L}} = 0.268 \text{ M}$$

5. A 10.0 g sample of sodium chloride is dissolved in enough water to make 0.5 L of solution. Find the molarity of the solution.

$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol}}{58.44 \text{ g}} = 0.171 \text{ mol} \times \frac{1 \text{ L}}{0.500 \text{ L}} = 0.342 \text{ M}$$

Unit 1 worksheet 3 Molar Concentration

IM Harris



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Unit 1worksheet 3 Molar Concentration Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Unit 1worksheet 3 Molar Concentration**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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