

Limiting and Excess Reactant, Theoretical and Percent Yield.

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8:51 pm

To find the excess reactant and limiting reactant, you can divide the amount of given moles with their coefficient in the balanced equation. The one that has a lower coefficient to mole ratio is the limiting reactant.

Example:



Another way is to compute their theoretical yield, the one that gives the smaller amount of product is the limiting reagent.

2 moles oxygen & 2 mol propane

$$\begin{aligned} & \frac{2 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{3 \text{ mol CO}_2}{1 \text{ mol C}_3\text{H}_8} = 6 \text{ mol} \\ & \frac{2 \text{ mol O}_2}{5 \text{ mol O}_2} \cdot \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} = 4.8 \text{ mol} \end{aligned}$$

How to find the percent yield.

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

If instead of 4.8 mol of carbon dioxide you got 4.5, what would be the percent yield?

$$\% \text{ yield} = \frac{4.5 \text{ mol}}{4.8 \text{ mol}} \times 100\% = 93.75\%$$

Then compute the error yield by subtracting the percent yield by 100.

$$100 - 93.75 = 6.25\%$$

How to convert ppm to molarity

$$\text{ppm} = \frac{\text{g solute}}{\text{g soln}} \times 10^6 \quad M = \frac{\text{mol of solute}}{\text{L of soln}}$$

example

A solution consisting of water and NaOH has a density of 1.05 g/mL and a solute concentration of 840 ppm. What is the molarity of this solution?

To convert grams of solution to L of solution we need density. As well as needing the Molar Mass to convert grams to moles.

$$\text{NaOH} = 39.998 = 40 \text{ g/mol}$$

Step by step process:

$$\text{ppm} = \frac{\text{g solute}}{\text{g solution}} \cdot 1 \times 10^{-6} \quad M = \frac{\text{mol of solute}}{\text{L of solution}}$$

1. First we will write down the ppm, to convert to the grams to moles.

$$\frac{840 \text{ g NaOH}}{1 \times 10^6 \text{ g soln}} \cdot \frac{1 \text{ mol}}{40 \text{ g NaOH}} = \frac{21 \text{ mol}}{1 \times 10^6 \text{ g soln}}$$

2. Then, we will multiply what we got to the substance's density to convert g/1x10⁶g to g/mL.

$$\frac{21 \text{ mol}}{1 \times 10^6 \text{ g}} \cdot \frac{1.05 \text{ g}}{\text{mol}} = 22$$

Unit 4 Limiting Reactant And Percent Yield Ws 7

RC Schank



Unit 4 Limiting Reactant And Percent Yield Ws 7:

Theoretical Chemical Engineering Abstracts ,1980 **Limiting Reactant and Percent** Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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