Name

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Chemistry - Unit 7 Reaction Equations Worksheet 1

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

- 1. $_1_C + _1_{H_2O} \rightarrow _1_{CO} + _1_{H_2}$
- 2. $_2MgO \rightarrow _2Mg + _1_O_2$
- _4_Al + _3_O₂ → _2_Al₂O₃

#3 Before A	fter
-------------	------

Before

- 4. _1_Zn + _1_H₂SO₄ → _1_ZnSO₄ + _1_H₂
- _1_Cl₂ + _2_Kl → _2_KCl + _1_l₂
- _2_CuCl → _2_Cu + _1_Cl₂
- _2_Na + _1_Cl₂ → _2_NaCl
- 8. _2_Al + _6_HCl → _2_AlCl₃ + 3 H₂
- 9. 2 Fe₂O₃ → 4 Fe + 3 O₂
- 10. $_4_P + _5_O_2 \rightarrow _2_P_2O_5$
- 11. $_1$ _Mg + $_2$ _HCl \rightarrow $_1$ _MgCl₂ + $_1$ _H₂
- 12. $_3H_2 + _1N_2 \rightarrow _2NH_3$
- 13. $_1$ _BaCl₂ + $_1$ _H₂SO₄ \rightarrow $_1$ _BaSO₄ + $_2$ _HCl
- 14. $_1$ _CH₄ + $_2$ _O₂ \rightarrow $_1$ _CO₂ + 2 H₂O

#14 Before After

- 15. a) ___ZnCl₂ + _ NH₄)₂S \rightarrow ___ZnS + ___ NH₄Cl
 - b) Find the molar mass of these reactants.
 - c) How many moles of ZnCl₂ would be in 25 g? How much mass would 0.55 moles of (NH₄)₂S have?

#8

After

Unit 7 Reaction Equations 1 Key

Clifford Lane

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Table of Contents Unit 7 Reaction Equations 1 Key

- 1. Understanding the eBook Unit 7 Reaction Equations 1 Key
 - The Rise of Digital Reading Unit 7 Reaction Equations 1 Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Unit 7 Reaction Equations 1 Key
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 7 Reaction Equations 1 Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Unit 7 Reaction Equations 1 Key
 - Personalized Recommendations
 - Unit 7 Reaction Equations 1 Key User Reviews and Ratings
 - Unit 7 Reaction Equations 1 Key and Bestseller Lists

- 5. Accessing Unit 7 Reaction Equations 1 Key Free and Paid eBooks
 - Unit 7 Reaction Equations 1 Key Public Domain eBooks
 - Unit 7 Reaction Equations 1 Key eBook Subscription Services
 - Unit 7 Reaction Equations 1 Key Budget-Friendly Options
- 6. Navigating Unit 7 Reaction Equations 1 Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Unit 7 Reaction Equations 1 Key Compatibility with Devices
 - Unit 7 Reaction Equations 1 Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 7 Reaction Equations 1 Key
 - Highlighting and Note-Taking Unit 7 Reaction Equations 1 Key
 - Interactive Elements Unit 7 Reaction Equations 1 Key
- 8. Staying Engaged with Unit 7 Reaction Equations 1 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 7 Reaction Equations 1 Key
- 9. Balancing eBooks and Physical Books Unit 7 Reaction Equations 1 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 7 Reaction Equations 1 Key
- 10. Overcoming Reading Challenges
 - $\circ\,$ Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 7 Reaction Equations 1 Key
 - Setting Reading Goals Unit 7 Reaction Equations 1 Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 7 Reaction Equations 1 Key
 - Fact-Checking eBook Content of Unit 7 Reaction Equations 1 Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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