

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this same solution had been added to the flask solution. This would be incorrect. **Just this solution.**
Then, calculate the molarities of the two solutions. The molarity of the solution is 1.5M. The molarity of the solution is 1.5M.
2. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molar concentration of this potassium chloride.

$$10.0 \text{ g} \times \frac{1 \text{ mol}}{74.55 \text{ g}} = 0.134 \text{ mol}$$

$$\text{Molarity} = \frac{0.134 \text{ mol}}{0.100 \text{ L}} = 1.34 \text{ M}$$

3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.25 L.

$$25.0 \text{ g} \times \frac{1 \text{ mol}}{110.98 \text{ g}} = 0.225 \text{ mol}$$

$$\text{Molarity} = \frac{0.225 \text{ mol}}{0.25 \text{ L}} = 0.90 \text{ M}$$

4. An unknown ionic solution's "molarity" is defined by preparing a 1.00 g of solution sample and measuring the volume of solution in mL. Calculate the molarity of the solution in the solution.

$$1.00 \text{ g} \times \frac{1 \text{ mol}}{100.0 \text{ g}} = 0.01 \text{ mol}$$

$$\text{Molarity} = \frac{0.01 \text{ mol}}{0.01 \text{ L}} = 1.00 \text{ M}$$

5. A 10.0 g solution sample is prepared by adding 10.0 g of CaCl_2 to enough water to make 0.25 L.

$$10.0 \text{ g} \times \frac{1 \text{ mol}}{110.98 \text{ g}} = 0.090 \text{ mol} \quad \text{Molarity} = \frac{0.090 \text{ mol}}{0.25 \text{ L}} = 0.36 \text{ M}$$

Unit 9 Worksheet 3 Molar Concentration Answer Key

A Gutmann



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