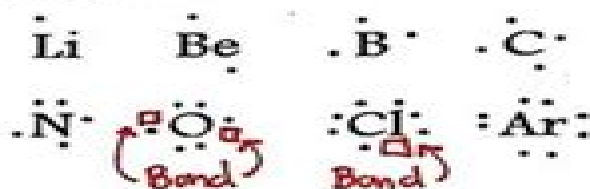


Lewis (Electron) Dot Diagrams:

Below you will find eight (8) electron dot diagrams for common elements. The dots around the element's symbol represent a certain type of electrons.



- What type of electrons possessed by an atom in its outermost energy level do the dots represent?

Valence electrons

- Why do you think these electrons in Carbon, Nitrogen, and Oxygen especially are spread out in the way that they are?

e^- repel until they have no choice

- If Chlorine has 7 dots, how many more can it possibly gain? 1 $7+1=8$
 - If Oxygen has 6 dots, how many more can it possibly gain? 2 $6+2=8$
 - What is the correlation between open electron spaces and possible bonds formed?

Direct - open space is a location for a bond

Practice:

1. Write the electron dot diagram for each of the following elements:

a. Hydrogen -
1 ve.



b. Barium -
2 ve.



c. Krypton -
8 ve.



d. Magnesium -
2 ve.



e. Oxygen -
6 ve.



f. Chlorine -
7 ve.



* Often, the .. will orient next to each other

2. Use electron dot diagrams to show the transfer of electrons in each of the following ionic bonds. Then, write the formula for each combination - Cation(s) first!

a. ^MPotassium and Chlorine



b. Beryllium and Fluorine



c. Rubidium and Oxygen



d. Calcium and Sulfur



Unit 5 Chemical Bonding Test Answers

George Facer



Unit 5 Chemical Bonding Test Answers:

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Table of Contents Unit 5 Chemical Bonding Test Answers

1. Understanding the eBook Unit 5 Chemical Bonding Test Answers
 - The Rise of Digital Reading Unit 5 Chemical Bonding Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 5 Chemical Bonding Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 5 Chemical Bonding Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 5 Chemical Bonding Test Answers
 - Personalized Recommendations
 - Unit 5 Chemical Bonding Test Answers User Reviews and Ratings
 - Unit 5 Chemical Bonding Test Answers and Bestseller Lists
5. Accessing Unit 5 Chemical Bonding Test Answers Free and Paid eBooks
 - Unit 5 Chemical Bonding Test Answers Public Domain eBooks
 - Unit 5 Chemical Bonding Test Answers eBook Subscription Services
 - Unit 5 Chemical Bonding Test Answers Budget-Friendly Options

6. Navigating Unit 5 Chemical Bonding Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 5 Chemical Bonding Test Answers Compatibility with Devices
 - Unit 5 Chemical Bonding Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 5 Chemical Bonding Test Answers
 - Highlighting and Note-Taking Unit 5 Chemical Bonding Test Answers
 - Interactive Elements Unit 5 Chemical Bonding Test Answers
8. Staying Engaged with Unit 5 Chemical Bonding Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 5 Chemical Bonding Test Answers
9. Balancing eBooks and Physical Books Unit 5 Chemical Bonding Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 5 Chemical Bonding Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 5 Chemical Bonding Test Answers
 - Setting Reading Goals Unit 5 Chemical Bonding Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 5 Chemical Bonding Test Answers
 - Fact-Checking eBook Content of Unit 5 Chemical Bonding Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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