

Name _____
Date _____

Unit 6: Stoichiometry – Intensive Review

Formula Mass & Percent Composition

1. The sum of the atomic masses of the atoms in one molecule of C_4H_{10} is called the

- (1) formula mass (3) percent abundance
(2) isotopic mass (4) percent composition

2. The gram-formula mass of a compound is 48 grams. The mass of 1.0 mole of this compound is

- (1) 1.0 g (2) 4.8 g (3) 48 g (4) 480 g

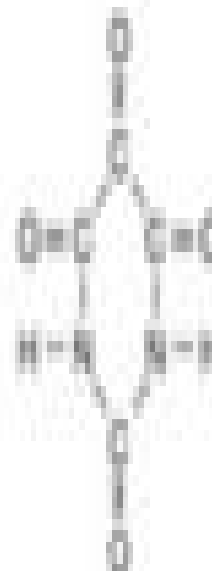
3. What is the gram-formula mass of $(NH_4)_3PO_4$?

- (1) 112 g/mol (2) 121 g/mol (3) 149 g/mol (4) 242 g/mol

4. What is the gram-formula mass of $Ca_3(PO_4)_2$?

- (1) 248 g/mol (2) 263 g/mol (3) 279 g/mol (4) 310 g/mol

14. Given the formula for a compound,



Which molecular and empirical formulas represent this compound?

- (1) $C_2H_2O_4$ and $C_2H_2O_4$ (3) $C_2H_2O_4$ and $C_2H_2O_4$
(2) $C_2H_2O_4$ and $C_2H_2O_4$ (4) $C_2H_2O_4$ and $C_2H_2O_4$

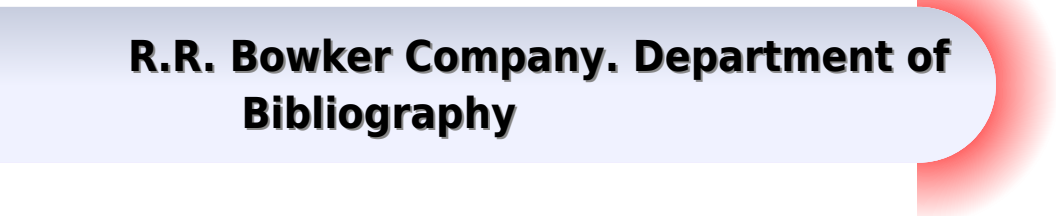
15. A compound has a molar mass of 90. grams per mole and the empirical formula CH_3O . What is the molecular formula of the compound?

- (1) CH_3O (2) $C_2H_6O_2$ (3) $C_3H_9O_3$ (4) $C_4H_{12}O_4$

16. The structural formula for methanol is CH_3 and the molecular

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