

CHEMISTRY SAMPLE TEST 1

$$PV = nRT \rightarrow n = \frac{PV}{RT}$$

$$n = 5,003$$

1. A sample of N_2 gas occupies 4.48 L volume under standard conditions (1 atm, $0^\circ C$). What is the mass of the sample? The atomic mass of N is 14 amu.

A: 5.6 g B: 56 g C: 2.8 g D: 28 g E: 22.4 g

2. Which properties are characteristic for the nonmetals?

1. high electrical conductivity
2. large ionization energy
3. high electronegativity
4. low electron affinity

A: 1,2

B: 2,3

C: 3,4

D: 2,4

E: 1,2,3

3. Concerning 1.2×10^{24} CO_2 molecules, which statements are true?

The molar mass is 44 g/mol.

1. it is 12 moles.
2. it occupies 1.2×22.4 L volume under standard conditions.
3. it has a mass of 88 grams.
4. it consists of 3.6×10^{24} atoms.

A: 1,3

B: 1,4

C: 2,3

D: 2,4

E: 3,4

$$\text{mole} = \frac{\text{particle}}{\text{Avogadro's no}}$$

gas 2 was molecules

4. Which of the following substances contain covalent bonds only?

1. $BaCl_2(aq)$

2. $CCl_4(l)$

3. $HCl(g)$

4. $NH_4Br(aq)$

A: 1,2,3,4

B: 1,2,4

C: 2,3

D: 1,3

E: 2,3,4

• covalent bond: two atoms share electrons
• ionic bond: one atom donates electron to the other
↑ ionic
1 non-metal + 1 metal

5. Which of the following atoms are isotopes of each other?

X: 11 protons, 11 electrons, 12 neutrons

Y: 11 protons, 10 electrons, 12 neutrons

Z: 11 protons, 11 electrons, 13 neutrons

W: 12 protons, 12 electrons, 12 neutrons

Same proton no
different neutron no

A: X and Y

B: X and W

C: Y and Z

D: Z and W

E: X and Z

6. An aqueous solution is prepared by dissolving 6 g NaOH in 250 mL final volume. What is the molar concentration of the solution?

The molar mass of NaOH is 40 g/mol

A: 6.4 mol/L

B: 1.6 mol/L

C: 16 mol/L

D: 0.16 mol/L

E: 64 mol/L

$$C_M = \frac{n}{V}$$

$$\frac{6 \text{ mol}}{250 / 1000 \text{ L}}$$

Unit Chemistry Test Test Number Answer Document

Rachel S Tattersall



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