

Gas Laws HOMEWORK

Lesson 1 Homework

1. Define pressure. *Force that a gas exerts on its container*

2. Convert the following:

a. 750 mmHg to atm

b. 550 mmHg to atm

c. 2.5 atm to torr

d. 2.1×10^3 Pa to atm

e. 30 Pa to atm

$$750 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.9868 \text{ atm}$$

$$550 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.72368 \text{ atm}$$

$$2.5 \text{ atm} \times (760 \text{ torr} / 1 \text{ atm}) = 1900 \text{ torr}$$

$$2.1 \times 10^3 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.02072 \text{ atm}$$

$$30 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 2.96 \times 10^{-4} \text{ atm}$$

3. Calculate the atmospheric pressure in kPa on a day when the pressure is 750 mm.

$$750 \text{ mm} \times (101.325 \text{ kPa} / 760 \text{ mm}) = 100.367 \text{ kPa} \rightarrow 100.4 \text{ kPa}$$

4. Calculate the pressure of a sample of gas (in mmHg) inside an open-end manometer if the mercury column is 35 mm higher on the side of the closed tube than on the atmosphere. The atmospheric pressure is 750 mm. Draw a picture to illustrate your answer.

$$750 \text{ mm} + (35 \text{ mmHg} / 1 \text{ mm}) = 785 \text{ mmHg}$$

750 mm mercury is 750 mm Hg because the atmospheric pressure, then it's 750 mmHg because 750 mmHg.

Since pressure doesn't change, you can leave it out of the equation. Then, since the atmospheric level is 750 mm.

$$785 \text{ mmHg} - 75 = 710 \text{ mmHg}$$

5. If a gas at a pressure of 150 mm in a 1.0 L container has its pressure increased to 400 mm, what is the new volume of the container? (Assume temperature is constant).

$$P_1 V_1 = P_2 V_2$$

$$(150 \text{ mm} \times 1.0 \text{ L}) = (400 \text{ mm} \times V_2)$$

$$V_2 = 0.375 \text{ L} \rightarrow 0.38 \text{ L}$$

6. What is the new pressure if a 5.00 closed container with a pressure of 2.5 atm is increased to 8.0 L?

$$P_1 V_1 = P_2 V_2$$

$$(2.5 \text{ atm} \times 5.00 \text{ L}) = (P_2 \times 8.0 \text{ L})$$

$$P_2 = 1.5625 \rightarrow 1.6 \text{ atm}$$

7. When 0.5 L of oxygen and 4.0 L of Nitrogen both at 0°C and 1.0 atm, are pumped into an evacuated 5.0 L rigid container, what is the final pressure in the container at 0°C? *Since temperature doesn't change, you can leave it out of the equation.*

So, final pressure of Nitrogen:

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 4.0 \text{ L}) = (P_2 \times 5.0 \text{ L})$$

$$P_2 = 0.8 \text{ atm}$$

So, final pressure of Oxygen:

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 0.5 \text{ L}) = (P_2 \times 5.0 \text{ L})$$

$$P_2 = 0.1 \text{ atm}$$

$$\text{Total Pressure} = P_1 + P_2 = P_3$$

$$1.0 + 0.8 + 0.1 = 1.9 \rightarrow 1.9 \text{ atm}$$

Assume both containers are 5.0 L each.

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 5.0 \text{ L}) = (P_2 \times 5.0 \text{ L})$$

$$P_2 = 1.0 \text{ atm} \rightarrow 1.0 \text{ atm}$$

Then, since both containers volume and both are directly proportional.

8. A sealed flask contains a gas with pressure P. If the number of moles of the gas is doubled, the absolute temperature is doubled and the volume of the container is quadrupled, what is the new pressure in terms of P?

Doubling the moles doubles the pressure.

$$P_1 V_1 = P_2 V_2$$

Quadrupling the volume cuts the pressure in 1/4.

$$P_1 V_1 = P_2 V_2$$

moles do no overall change in both P & V

9. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change within the container?

$$P_1 V_1 = P_2 V_2$$

Going from 1 mole to 5 moles results in an increase of pressure by 4 times.

$$P_1 V_1 = P_2 V_2$$

10. If a balloon containing 1000 L of gas at 40.°C and 760 mmHg rises to an altitude where the pressure is 500 mmHg and the temperature is 10.0°C, what is the new volume of the balloon (in L)?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \rightarrow V_2 = ?$$

$$(760 \text{ mmHg} \times 1000 \text{ L}) / (40 + 273) = (500 \text{ mmHg} \times V_2) / (10 + 273)$$

$$V_2 = 2017.76 \rightarrow 2000 \text{ L}$$

11. A sample of nitrogen gas has a pressure of 87.5 mmHg in a 500. mL flask. What is the pressure of this gas sample when it is transferred to a 225.0 mL flask at the same temperature?

$$P_1 V_1 = P_2 V_2$$

$$(87.5 \text{ mmHg} \times 500 \text{ mL}) = (P_2 \times 225 \text{ mL})$$

$$P_2 = 194.44 \rightarrow 194 \text{ mmHg}$$

12. You have 4.5 L of NO at a temperature of 22.0°C. What volume would the NO occupy at 47°C? (Assume the pressure is constant).

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \rightarrow V_2 = ?$$

$$(4.5 \text{ L} \times (22.0 + 273)) = (V_2 \times (47 + 273))$$

$$V_2 = 4.00 \text{ L} \rightarrow 4.0 \text{ L}$$

13. You have a sample of a gas in a flask with a volume of 240 mL. At 35.0°C the pressure of the gas is 540 mmHg. If you decrease the temperature to -5.0°C, what is the gas pressure at the lower temperature? *Since volume doesn't change, you can leave it out of the equation.*

$$P_1 / T_1 = P_2 / T_2$$

$$(540 \text{ mmHg} \times (35 + 273)) = (P_2 \times (-5 + 273))$$

$$P_2 = 400.48 \rightarrow 400 \text{ mmHg}$$

Unit Gas Laws Homework Chemistry Answers

N Noddings



Unit Gas Laws Homework Chemistry Answers:

Introductory Chemistry Problems: The Chemistry And Physics Behind The Numbers Michael E Green, 2025-07-18 The essential point of the book is to show students in Introductory Chemistry classes how to obtain answers that are at least reasonable and to realize when there must have been an error because the answer falls outside the range of possible values. Students often leave ridiculous answers because they simply plug into equations they have memorized correctly or incorrectly. They have not learned to think about the meaning of the numbers that appear in their calculations. This book attempts to show students how to correct this.

Homework Helpers: Chemistry Greg Curran, 2025-09-12 Homework Helpers Chemistry is a user friendly review book that will make every student or parent trying to help their child feel like he or she has a private Chemistry tutor. Concepts are explained in clear easy to understand language and problems are worked out with step by step methods that are easy to follow. Each lesson comes with numerous review questions and answer keynotes that explain each correct answer and why it is correct. This book covers all of the topics in a typical one year Chemistry curriculum including A systematic approach to problem solving conversions and the use of units Naming compounds writing formulas and balancing chemical equations Gas laws chemical kinetics acids and bases electrochemistry and more. While Homework Helpers Chemistry is an excellent review for any standardized Chemistry test including the SAT II its real value is in providing support and guidance during the year's entire course of study.

Chemistry John Kenkel, Paul B. Kelter, David S. Hage, 2000-09-21 What a great idea an introductory chemistry text that connects students to the workplace of practicing chemists and chemical technicians. Tying chemistry fundamentals to the reality of industrial life. Chemistry An Industry Based Introduction with CD ROM covers all the basic principles of chemistry including formulas and names chemical bonds.

Chemistry Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics. Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases. The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry. The 100 Series science books span grades 5 to 12. The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science. The books include engaging grade appropriate activities and clear thumbnail answer keys. Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics. The series will be aligned to current science standards.

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards **Holt Chemistry** Ralph Thomas Myers,2004 **Novice Teacher Action** Anne Liu Kern,2007 **The Role of Language in Content Pedagogy** Lay Hoon Seah,Rita Elaine Silver,Mark Charles Baildon,2022-11-01 This book explores the importance of language in content learning It focuses on teachers roles knowledge and understanding of language in school contexts including academic language and disciplinary languages to support students It examines teachers language related knowledge base for content teaching which include teachers knowledge of and about language knowledge of their students and their pedagogical knowledge This book also explores how teachers knowledge of language students and content are linked as part of a larger pedagogical content knowledge which includes knowledge of the role of language in content learning As well it further considers literacy and literacies as part of this examination of teachers knowledge of language

Basic Chemistry Concepts and Exercises John Kenkel,2010-10-04 Chemistry can be a daunting subject for the uninitiated and all too often introductory textbooks do little to make students feel at ease with the complex subject matter Basic Chemistry Concepts and Exercises brings the wisdom of John Kenkel s more than 35 years of teaching experience to communicate the fundamentals of chemistry in a practical d **Illinois Chemistry Teacher** ,1998 Chemical Principles Richard Earl Dickerson,1984 The Software Encyclopedia 2001 ,2001 *The Educational Value of Chemical Demonstrations in the College Prep Chemistry Classroom* Katherine E. Hagerman,2010 *Advancing the STEM Agenda* Cindy P. Veenstra,Fernando F. Padró,Julie A. Furst-Bowe,2012-05-15 In July 2011 the ASQ Education Division held its first Advancing the STEM Science Technology Engineering and Mathematics Agenda in Education the Workplace and Society Conference at the University of Wisconsin Stout This publication is a selection of papers and workshops from this groundbreaking conference The ideas presented here will help other educators and policy makers to develop their own innovative high impact ideas for inspiring student interest in STEM careers improving the delivery of STEM education at their schools and colleges and helping STEM college graduates transition to the workplace The chapters in this book reflect research and best practices integrating the ideas of continuous improvement in combination with a can do attitude to provide a valuable resource that will lead others to consider similar innovative and collaborative educational structures that will drive more interest in STEM majors in college and provide for our next generation of scientists technicians and engineers Prior to reviewing Advancing the STEM Agenda I had a list in my mind of topics that I hoped would be addressed I m very pleased with how many are covered and covered well This project succeeds at the challenge of providing not only beneficial breadth but also important depth Because our public private partnership has been committed explicitly to continuous improvement for more than a decade I couldn t help but notice as the editors also point out in their conclusion the extent to which

continuous improvement is a common thread throughout the book That speaks to the book s practical utility in many settings and on a long term basis No less valuable is the discussion of student motivation by many of the authors which STEM teachers in our area have identified as a major issue of interest to them in recent surveys Richard Bogovich Executive Director Rochester Area Math Science Partnership Minnesota Veenstra Padr and Furst Bowe provide a huge contribution to the field of STEM education We all know the statistics and of the huge need in the area of STEM students and education but what has been missing are application and success stories backed by research and modeling The editors have successfully contributed to our need by focusing on collaborative models building the K 12 pipeline showing what works at the collegiate level connecting across gender issues and illustrating workforce and innovative ideas John J Jasinski President Northwest Missouri State University Advancing the STEM Agenda provides a broad set of current perspectives that will contribute in many ways to advancing the understanding and enhancement of education in science education and engineering This work is packed with insights and perspectives from experienced educators and bridges the transition from education to workplace John Dew Senior Vice Chancellor Troy University *Foundations of College Chemistry* Morris Hein,Susan Arena,2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **Physical Chemistry Using MathCAD** Joseph H. Noggle,1997 Mathcad R MathSoft Inc is a computer program for mathematics that can do not only calculations but symbolic algebra calculus differential equations other advanced mathematical techniques Its advantage over competing programs is its ability to keep track of units do unit conversions its ease of learning use This book is designed to teach the reader how to use the program in the context of learning physical chemistry with examples from thermodynamics kinetics transport processes quantum mechanics While it is primarily intended for students it will also be useful for graduate scientists engineers who wish to review the subject or to learn about new methods of doing scientific engineering calculations using a microcomputer To order Pike Creek Publishing Company 32 Donegal Court Newark DE 19711 302 234 3320 *Chemistry* Zumdahl,Steven S. Zumdahl,2003-07 Contains chapter discussions key term definitions and practice chapter tests **Nucleation of Particles from the Gas Phase** , Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall,Donald J. DeCoste,2000

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