

Classifying and Balancing Equations
Multiple Choice

Name _____
Date _____

Per _____

PRACTICE TEST

- _____ 1. A chemical reaction is a process in which
a. products change into reactants. c. substances can change state
b. the law of conservation of mass applies d. all of these
- _____ 2. During a chemical reaction,
a. new elements are produced c. atoms are destroyed
b. atoms are rearranged d. elements are destroyed
- _____ 3. An equation is balanced by
a. changing subscripts c. erasing elements as necessary
b. adding coefficients d. adding elements as necessary
- _____ 4. An atom's ability to undergo chemical reactions is determined by its
a. protons b. innermost electrons c. neutrons d. valence electrons
- _____ 5. What are the reactants in the following chemical equation:
$$\text{Zn} + \text{CuSO}_4 \longrightarrow \text{ZnSO}_4 + \text{Cu}$$

a. zinc and copper c. zinc and copper (II) sulfate
b. zinc sulfate and copper d. only zinc
- _____ 6. What are the products in the above equation?
a. zinc and copper c. zinc and copper (II) sulfate
b. zinc sulfate and copper d. only zinc

Short Answer

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.

- _____ 7. $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$
- _____ 8. $\text{H}_2\text{CO}_3 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- _____ 9. $\text{CaCO}_3 + \text{HCl} \longrightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
- _____ 10. $\text{AgNO}_3 + \text{Zn} \longrightarrow \text{Zn}(\text{NO}_3)_2 + \text{Ag}$
- _____ 11. $\text{C}_3\text{H}_8 + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- _____ 12. $\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$

Write a balanced equation for each of the following reactions:

13. Magnesium chloride is the product of a reaction between magnesium and chlorine.

14. Copper (II) hydroxide and potassium sulfate are produced when potassium hydroxide reacts with copper (II) sulfate.

Unit 6 Practice Test - Key

1. b
2. b
3. b
4. d
5. c
6. B
7. Single replacement 1,1,1,1
8. Decomposition 1,1,1
9. Double replacement 1,2,1,1-
CaCO₃ is ppt (see Ref. Table)
10. Single replacement 2,1,1,2

11. Combustion 1, 5, 3, 4
12. Combustion 1,3,2,3
13. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$ 1,1,1
14. $\text{KOH} + \text{CuSO}_4 \rightarrow \text{Cu}(\text{OH})_2 + \text{K}_2\text{SO}_4$ 2,1,1,1

Unit Test Review Chemical Reactions Answer Key

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The Electro-chemist and Metallurgist and Metallurgical Review ,1902 Fossil Energy Update ,1980 Energy Research Abstracts ,1995 Solar Energy Update ,1981 **Heath Science Connections 9** Jack L. Candido,Edward S. James,Ronald E. Phillips,1987 Resources for Teaching Middle School Science Smithsonian Institution,National Academy of Engineering,National Science Resources Center of the National Academy of Sciences,Institute of Medicine,1998-04-30

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Table of Contents Unit Test Review Chemical Reactions Answer Key

1. Understanding the eBook Unit Test Review Chemical Reactions Answer Key
 - The Rise of Digital Reading Unit Test Review Chemical Reactions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit Test Review Chemical Reactions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit Test Review Chemical Reactions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit Test Review Chemical Reactions Answer Key
 - Personalized Recommendations
 - Unit Test Review Chemical Reactions Answer Key User Reviews and Ratings
 - Unit Test Review Chemical Reactions Answer Key and Bestseller Lists
5. Accessing Unit Test Review Chemical Reactions Answer Key Free and Paid eBooks
 - Unit Test Review Chemical Reactions Answer Key Public Domain eBooks
 - Unit Test Review Chemical Reactions Answer Key eBook Subscription Services
 - Unit Test Review Chemical Reactions Answer Key Budget-Friendly Options

6. Navigating Unit Test Review Chemical Reactions Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit Test Review Chemical Reactions Answer Key Compatibility with Devices
 - Unit Test Review Chemical Reactions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit Test Review Chemical Reactions Answer Key
 - Highlighting and Note-Taking Unit Test Review Chemical Reactions Answer Key
 - Interactive Elements Unit Test Review Chemical Reactions Answer Key
8. Staying Engaged with Unit Test Review Chemical Reactions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit Test Review Chemical Reactions Answer Key
9. Balancing eBooks and Physical Books Unit Test Review Chemical Reactions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit Test Review Chemical Reactions Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit Test Review Chemical Reactions Answer Key
 - Setting Reading Goals Unit Test Review Chemical Reactions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit Test Review Chemical Reactions Answer Key
 - Fact-Checking eBook Content of Unit Test Review Chemical Reactions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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