Chapter 16 Redox chemistry and corrosion

QL.

Identify each of the following half equations as involving either oxidation or reduction:

- $n = Nn(s) \rightarrow Nn^*(nq) + e^-$
- b Cl₂(g) + 2e⁻ → 2CΓ (aq)
- $c = S(s) + 2e \rightarrow S^{2}(aq)$
- d $Zn(s) \rightarrow Zn^{2+}(aq) + 2e$

A.L.

- a oxidation
- b reduction
- e reduction
- d oxidation

Q2.

Balance the following half equations and then identify each as representing either an oxidation or a reduction reaction:

- a $Fe(s) \rightarrow Fe^{3s}(aq)$
- b $K(s) \rightarrow K^*(aq)$
- $e \to F_2(g) \to F_2(aq)$
- d $O_2(g) \rightarrow O^2$ (aq)

A.2.

- a Fe(s) → Fe^{3*}(aq) + 3e⁻ oxidation
- b K(s) → K⁺(aq) + e⁻ oxidation
- e $F_2(g) + 2e^- \rightarrow 2F^-(aq)$ reduction
- d $O_2(g) + 4e^- \rightarrow 2O^2$ (aq) reduction

Q3.

Iron reacts with hydrochloric acid according to the ionic equation:

 $Fe(s) + 2H^*(aq) \rightarrow Fe^{2*}(aq) + H_2(g)$

- a What has been oxidised in this reaction? What is the product?
- b Write a half equation for the oxidation reaction.
- Identify the oxidant.
- d What has been reduced in this reaction? What is the product?
- Write a half equation for the reduction reaction.
- Identify the reductant.

A.3.

- a Fe(s) has been oxidised to Fe^{2*}(aq)
- b $Fe(s) \rightarrow Fe^{2s}(aq) + 2e$
- c H*(aq)
- d H*(aq) has been reduced to H₂
- e $2H^*(aq) + 2e \rightarrow H_2(g)$
- f Fe(s)

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Electrochemistry and Corrosion Science Nestor Perez, 2016-09-13 The second edition of this textbook includes refined text in each chapter new sections on corrosion of steel reinforced concrete and on cathodic protection of steel reinforced bars embedded in concrete and some new solved examples The book introduces mathematical and engineering approximation schemes for describing the thermodynamics and kinetics of electrochemical systems which are the essence of corrosion science in addition to electrochemical corrosion forms of corrosion and mechanisms of corrosion This approach should capture the reader s attention on the complexity of corrosion Thus the principles of electrochemistry and electrochemical cells are subsequently characterized in simple electrolytes from a thermodynamics point of view

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Table of Contents Worked Solutions Heinemann Chemistry 1

- 1. Understanding the eBook Worked Solutions Heinemann Chemistry 1
 - The Rise of Digital Reading Worked Solutions Heinemann Chemistry 1
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Worked Solutions Heinemann Chemistry 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Worked Solutions Heinemann Chemistry 1
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Worked Solutions Heinemann Chemistry 1

- Personalized Recommendations
- Worked Solutions Heinemann Chemistry 1 User Reviews and Ratings
- Worked Solutions Heinemann Chemistry 1 and Bestseller Lists
- 5. Accessing Worked Solutions Heinemann Chemistry 1 Free and Paid eBooks
 - Worked Solutions Heinemann Chemistry 1 Public Domain eBooks
 - Worked Solutions Heinemann Chemistry 1 eBook Subscription Services
 - Worked Solutions Heinemann Chemistry 1 Budget-Friendly Options
- 6. Navigating Worked Solutions Heinemann Chemistry 1 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Worked Solutions Heinemann Chemistry 1 Compatibility with Devices
 - Worked Solutions Heinemann Chemistry 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Worked Solutions Heinemann Chemistry 1
 - Highlighting and Note-Taking Worked Solutions Heinemann Chemistry 1
 - Interactive Elements Worked Solutions Heinemann Chemistry 1
- 8. Staying Engaged with Worked Solutions Heinemann Chemistry 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Worked Solutions Heinemann Chemistry 1
- 9. Balancing eBooks and Physical Books Worked Solutions Heinemann Chemistry 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Worked Solutions Heinemann Chemistry 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Worked Solutions Heinemann Chemistry 1
 - Setting Reading Goals Worked Solutions Heinemann Chemistry 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Worked Solutions Heinemann Chemistry 1

- Fact-Checking eBook Content of Worked Solutions Heinemann Chemistry 1
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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