

Write the formulas for each of the following covalent compounds.

1. sulfur tetroxide _____
2. phosphorus trifluoride _____
3. iodine monobromide _____
4. silicon tetrahydride _____
5. selenium dioxide _____
6. oxygen dichloride _____
7. bromine pentachloride _____
8. oxygen difluoride _____
9. phosphorus pentachloride _____
10. nitrogen monoxide _____
11. dinitrogen tetroxide _____
12. diphosphorus pentoxide _____
13. disulfur difluoride _____
14. tetraphosphorus trisulfide _____
15. dinitrogen tetrachloride _____

Name each of the following covalent compounds.

1. CO_2 _____
2. SO_3 _____
3. N_2O_3 _____
4. N_2O_5 _____
5. PCl_3 _____
6. NH_3 _____

Worksheet Molecular Compounds Chemistry A Study Of Matter 6 Key

Emilie Sanchez



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Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet **Chemistry, Student Solutions Manual** Neil D. Jespersen,James E. Brady,Alison Hyslop,2011-03-15 Chemistry The Study of Matter and Its Changes Sixth Edition will provide the necessary practice support and individualised instruction that ensures success in the General Chemistry course This text provides the forum for problem solving and concept mastery of chemical phenomena that leads to proficiency and success in the General Chemistry course This edition will continue a molecular basis of chemistry tradition but in a manner that overtly and repeatedly reinforces the way properties at the molecular level are related to properties we observe at the macroscopic level The unique chemical tools approach employed in this book provides a way of thinking that helps those students develop the ability to analyse and solve both mathematical and conceptual problems This text follows the successful three step approach described as ANALYSIS SOLUTION and IS THE ANSWER REASONABLE This encourages the student to think about the problem before attempting to solve it then working through the solution and finally asking the important question Does the answer make sense There are problem sets called Bringing It Together that contain problems which require students to bring together concepts from two or more of the preceding chapters This reinforces learned concepts and builds concept mastery

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