

Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations - Set 2

Q-1: Balance the following chemical equations:

- i) $\text{CS}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{SO}_2$
- ii) $\text{SO}_2 + \text{H}_2\text{S} \rightarrow \text{H}_2\text{O} + \text{S}$
- iii) $\text{CaC}_2 + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$
- iv) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
- v) $\text{HgO} + \text{heat} \rightarrow \text{Hg} + \text{O}_2$

Q-2: Fill in the blanks

- i) Respiration involves the oxidation of _____ in the body.
- ii) When water is electrolytically decomposed, H_2 and O_2 gas are obtained at _____ and _____ respectively.
- iii) When H_2 gas is passed over heated copper(II) oxide, the residue left behind is that of _____.
- iv) Aluminium is highly _____ to corrosion.
- v) The reactions that involve absorption of heat are called _____.

Q-3: Salt formation occurs when small amount of acid is added to

- i) water
- ii) acid
- iii) base
- iv) pure water

Q-4: A visually impaired student can use which of the following as an acid-base indicator?

- i) Litmus
- ii) Red cabbage
- iii) Onion
- iv) Petunia Leaves

Q-5: Match the column I with appropriate entries of Column II

Column I	Column II
A) Silver nitrate + Sodium chloride	i) Red
B) Lead nitrate + Potassium iodide	ii) Black
C) Sodium sulphate + Barium chloride	iii) Yellow

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