

Unit 4 - Lesson 2 - Exploration 1 Notes

- A balanced chemical equation demonstrates the conservation of matter. So, you can use an equation to make predictions about the amounts of matter a reaction should produce.

Quantifying Moles in Chemical Reactions

- Atoms and molecules are too small to be measured individually, so chemists use a unit of measurement called the mole to quantify matter in chemical reactions.
- The mole is referred to as "the chemist's dozen" because it can be used to compare different substances by count.
- Atoms and molecules are much, much smaller than everyday items, such as eggs, so the number of particles in one mole of a substance is incredibly large.
- This number, which is approximately 6.02×10^{23} , is also called Avogadro's number.
- It was originally defined as the number of atoms in 12 g of carbon-12.
- You can also use dimensional analysis to predict how much matter a chemical reaction should produce.
- You can use the coefficients in this equation to make predictions about amounts of reactants and products. To do this, you must use a conversion factor that shows the coefficients as mole quantities.
- A conversion factor is a ratio between two quantities that are written as a fraction.
- By definition, a conversion factor always equals one, so either quantity can be written as the numerator.
- These conversion factors have the correct mole quantities based on the balanced chemical equation. The arrangement of the numerator and denominator depends on the specific problem you are solving.

Quantifying Mass in Chemical Reactions

- Using a balanced chemical equation to calculate quantities of matter, such as moles, volumes, or masses, is called *stoichiometry*.
- Stoichiometry allows you to convert from a quantity that is difficult to measure, such as the number of atoms in a substance, to a quantity that is easy to measure, such as the mass in grams.
- According to the table, the mass of one mole differs from one substance to another. However, as you saw in the earlier table, the number of particles in one mole does not differ by the substance.
- The mass in grams of one mole of an element or compound is called its molar mass.
- The molar mass of a single monatomic element, such as carbon, is equivalent to its atomic mass, as listed on the periodic table, expressed in grams. So, the mole is a unit that allows you to convert from grams, which you can measure on a balance, to individual molecules or atoms.
- The molar mass of a substance is expressed in grams per mole, or g/mol.

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