

History of the Atom



Democritus

He proposed that all matter, including space and time, was made up of small units named atoms. He did no experiments and had little evidence, but his idea was kept on by Lucretius



Mendeleev

Created the periodic table of elements which organizes elements according to their similarities. His Period Law states that "physical and chemical of the elements are periodic functions of their atomic numbers."



Henri Becquerel

Discovered radioactivity, which earned him a Nobel Prize. Also discovered that rays emitted from uranium caused gases to ionize



Ernest Rutherford

Used the gold foil experiment to discover the modern model of an atom. He concluded that all positive charges were centralized, while negative electrons orbited the nucleus



Niels Bohr

Explained that outer orbits in an atom could hold more electrons than the inner orbits. By knowing this, one can determine the atom's chemical properties. He also gave birth to the idea that electrons emit light by jumping orbits.



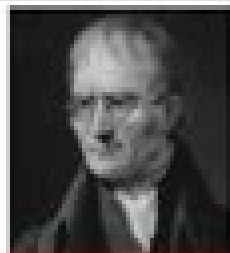
Heisenberg

Thought of the Uncertainty Principle. This principle states that one can never know the exact location and energy of an electron simultaneously.



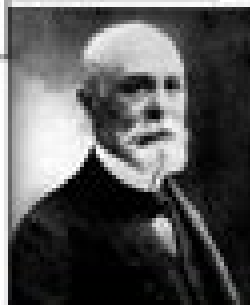
Dalton

Agreed that all matter was created by atoms, which he believed were indestructible. He also stated that compounds are created by combining two atoms, and that all atoms of given elements are identical in their mass and properties.



Eugene Goldstein

Discovered positive particles. He noted that the particles had a charge equal and opposite to the electron.



J.J. Thomson

Discovered the electron. He experimented by testing and studying the nature of electric discharge in a high vacuum cathode-ray tube.



Millikan

Used an oil drop experiment to determine the charge of a single electron.



Frederick Soddy

Came up with the concept of isotopes. Explain with Ernest Rutherford that radiation is due to transmutation of elements.



Chadwick

Discovered the neutron. The neutron helps balance out protons in the nucleus of an atom.



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Mark Steyvers



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