Name

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# Chemistry - Unit 7 Reaction Equations Worksheet 1

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

- 1.  $_1_C + _1_{H_2O} \rightarrow _1_{CO} + _1_{H_2}$
- 2.  $_2MgO \rightarrow _2Mg + _1_O_2$
- \_4\_Al + \_3\_O<sub>2</sub> → \_2\_Al<sub>2</sub>O<sub>3</sub>

#3 Before A	fter
-------------	------

Before

- 4. \_1\_Zn + \_1\_H<sub>2</sub>SO<sub>4</sub> → \_1\_ZnSO<sub>4</sub> + \_1\_H<sub>2</sub>
- \_1\_Cl<sub>2</sub> + \_2\_Kl → \_2\_KCl + \_1\_l<sub>2</sub>
- \_2\_CuCl → \_2\_Cu + \_1\_Cl<sub>2</sub>
- \_2\_Na + \_1\_Cl<sub>2</sub> → \_2\_NaCl
- 8. \_2\_Al + \_6\_HCl → \_2\_AlCl<sub>3</sub> + 3 H<sub>2</sub>
- 9. 2 Fe<sub>2</sub>O<sub>3</sub> → 4 Fe + 3 O<sub>2</sub>
- 10.  $_4_P + _5_O_2 \rightarrow _2_P_2O_5$
- 11.  $_1$ \_Mg +  $_2$ \_HCl  $\rightarrow$   $_1$ \_MgCl<sub>2</sub> +  $_1$ \_H<sub>2</sub>
- 12.  $_3H_2 + _1N_2 \rightarrow _2NH_3$
- 13.  $_1$ \_BaCl<sub>2</sub> +  $_1$ \_H<sub>2</sub>SO<sub>4</sub>  $\rightarrow$   $_1$ \_BaSO<sub>4</sub> +  $_2$ \_HCl
- 14.  $_1$ \_CH<sub>4</sub> +  $_2$ \_O<sub>2</sub>  $\rightarrow$   $_1$ \_CO<sub>2</sub> + 2 H<sub>2</sub>O

#### #14 Before After

- 15. a)  $\_\_ZnCl_2 + \_NH_4)_2S \rightarrow \_\_ZnS + \_\_NH_4Cl$ 
  - b) Find the molar mass of these reactants.
  - c) How many moles of ZnCl<sub>2</sub> would be in 25 g? How much mass would 0.55 moles of (NH<sub>4</sub>)<sub>2</sub>S have?

#8

After

# **Unit 7 Reaction Equations 1 Key**

J. R. Cobble, C. E. Jordan, David Arthur Rice

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