

Unit 10 Stoichiometry Test - Short Answer

Answer all questions including units and balanced equations where appropriate.

1. How many moles of aluminium are produced when completely with 2.5 mol of FeCl_3 ?

$$2\text{Al} + 3\text{FeCl}_3 \rightarrow 3\text{Fe} + \text{Al}_2\text{Cl}_6$$
2. How many grams of FeCl_3 are produced when 2.5 moles of aluminium reacts completely with FeCl_3 ?

$$2\text{Al} + 3\text{FeCl}_3 \rightarrow 3\text{Fe} + \text{Al}_2\text{Cl}_6$$
3. The reaction below shows that decomposition of FeCl_3 is complete. How many grams of oxygen are produced when 20.7g FeCl_3 is heated?

$$2\text{FeCl}_3 \rightarrow 2\text{Fe} + 3\text{Cl}_2$$
4. A reaction that has been calculated to produce 2.0g of AgCl actually produces 2.2g of AgCl . What is the percent yield?
5. How many moles of silver chloride are produced when 20.0g of silver reacts with 20.0g of hydrochloric acid according to this equation.

$$\text{HCl} + \text{Ag} \rightarrow \text{AgCl} + \text{H}_2$$

6. Percent Yield is

7. Amount of 20.0g produced is

8. How actually yields 20.0g of AgCl , what is the percent yield?

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