

ATOMS & THE PERIODIC TABLE

whit test

ATOMS & THE PERIODIC TABLE UNIT TEST

Name: _____

Complete the sentences.

1. The order of elements in the periodic table is based on the number of _____ in the nucleus.

2. Because atoms of elements in the same group of the periodic table have the same _____, they have similar properties.

3. Neon is an inert gas because its outer _____ is full of electrons.

10. Explain how a neutron defines a nuclide, and explain the difference between the nuclides _____.

11. How many electrons does Boron have (Be must they have)?

A. 4
B. 5
C. 6
D. 10

12. Elements in the same group on the periodic table have similar chemical properties.

A. True
B. False

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13. The number of protons in an atom is equal to the atomic number.

A. True
B. False

14. Based on the Bohr model of Boron, how many protons will it most likely have?

A. 4
B. 5
C. 6
D. 10

15. How is atomic mass calculated?

A. By combining the total count of protons and electrons in an atom.
B. By adding up the masses of protons, neutrons, and electrons in an atom.
C. By dividing the count of protons by the count of neutrons.
D. By multiplying the mass of electrons by the count of protons.

16. Elements are made up of different types of atoms.

A. True
B. False

17. How many electrons are needed to fill the outer shell (the Valence)?

A. 2
B. 4
C. 6
D. 8

18. What is an isotope?

A. An isotope is a form of an element with a positive charge.
B. An isotope is a form of an element with a different number of protons.
C. An isotope is an atom with a different number of protons than electrons.
D. An isotope is a variant of an element with the same number of protons but a different number of neutrons.

19. How many valence electrons does Boron have?

A. 4
B. 5
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20. What does it mean to say that some elements are reactive and others are nearly unreactive?

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Helen Dolman MacCracken



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