

CHEMISTRY SAMPLE TEST 1

3,0822

$n = 5,003$

$q = nF$

1. A sample of N_2 gas occupies 4.48 L volume under standard conditions (1 atm, $0^\circ C$). What is the mass of the sample? The atomic mass of N is 14 amu.

A: 5.6 g B: 56 g C: 2.8 g D: 28 g E: 22.4 g

2. Which properties are characteristic for the nonmetals?

- ~~1~~ high electrical conductivity
~~2~~ large ionization energy
~~3~~ high electronegativity
~~4~~ low electron affinity

~~A: 1,2~~

B: 2,3

C: 3,4

D: 2,4

~~E: 1,2,3~~

3. Concerning 1.2×10^{24} CO_2 molecules, which statements are true?

1. it is 12 moles.
 2. it occupies 1.2×22.4 L volume under standard conditions.
 3. it has a mass of 88 grams.
 4. it consists of 3.6×10^{24} atoms.

$\text{mole} = \frac{\text{particle}}{\text{Avogadro's no}}$

A: 1,3 B: 1,4 C: 2,3 D: 2,4 E: 3,4

4. Which of the following substances contain covalent bonds only?

- ~~1~~ $BaCl_2(aq)$ 2. $CCl_4(l)$ 3. $HCl(g)$ 4. $NH_4Br(aq)$

A: 1,2,3,4 B: 1,2,4 C: 2,3 D: 1,3 E: 2,3,4

• ionic bond: 1 metal + 2 non-metals
 • covalent bond: 2 non-metals
 • ionic bond: 1 metal + 1 non-metal
 • ionic bond: 1 metal + 2 non-metals
 • ionic bond: 1 metal + 1 non-metal

5. Which of the following atoms are isotopes of each other?

- ~~X~~: 11 protons, 11 electrons, 12 neutrons
 Y: 11 protons, 10 electrons, 12 neutrons
 Z: 11 protons, 11 electrons, 13 neutrons
 W: 12 protons, 12 electrons, 12 neutrons

Same proton no
 Different neutron no

A: X and Y B: X and W C: Y and Z D: Z and W E: X and Z

6. An aqueous solution is prepared by dissolving 6 g NaOH in 250 mL final volume. What is the molar concentration of the solution?

The molar mass of NaOH is 40 g/mol

- A: 6.4 mol/L
 B: 1.6 mol/L
 C: 16 mol/L
 D: 0.16 mol/L
 E: 64 mol/L

$C_m = \frac{n}{V}$

$\frac{6 \text{ g NaOH}}{40 \text{ g/mol}} = 0.15 \text{ mol}$
 $\frac{0.15 \text{ mol}}{0.250 \text{ L}} = 0.6 \text{ mol/L}$

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SJ Ball



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1. Understanding the eBook Unit Chemistry Test Test Number Answer Document
 - The Rise of Digital Reading Unit Chemistry Test Test Number Answer Document
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit Chemistry Test Test Number Answer Document
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit Chemistry Test Test Number Answer Document
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit Chemistry Test Test Number Answer Document
 - Personalized Recommendations
 - Unit Chemistry Test Test Number Answer Document User Reviews and Ratings
 - Unit Chemistry Test Test Number Answer Document and Bestseller Lists
5. Accessing Unit Chemistry Test Test Number Answer Document Free and Paid eBooks

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 - Unit Chemistry Test Test Number Answer Document Enhanced eBook Features
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 - Highlighting and Note-Taking Unit Chemistry Test Test Number Answer Document
 - Interactive Elements Unit Chemistry Test Test Number Answer Document
- 8. Staying Engaged with Unit Chemistry Test Test Number Answer Document
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit Chemistry Test Test Number Answer Document
- 9. Balancing eBooks and Physical Books Unit Chemistry Test Test Number Answer Document
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit Chemistry Test Test Number Answer Document
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit Chemistry Test Test Number Answer Document
 - Setting Reading Goals Unit Chemistry Test Test Number Answer Document
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit Chemistry Test Test Number Answer Document
 - Fact-Checking eBook Content of Unit Chemistry Test Test Number Answer Document
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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