

Worksheet 1 – Scientific Notation and Significant Figures

1. Convert each of the following into scientific notation.

- | | | |
|---------------------------|-------------------------|--------------------------|
| a) 3427 | b) 0.00456 | c) 123,453 |
| d) 172 | e) 0.000984 | f) 0.502 |
| g) 3100.0×10^2 | h) 0.0114×10^4 | i) 107.2 |
| j) 0.0000455 | k) 2205.2 | l) 30.0×10^{-2} |
| m) 0.982×10^{-3} | n) 0.0473 | o) 650.502 |
| p) 3.03×10^{-1} | q) 20.4×10^5 | r) 1.29 |
| s) 0.00565 | t) 1362205.2 | u) 450.0×10^3 |
| v) 1000×10^{-3} | | |

2. Determine the number of significant figures in each of the following:

- | | | |
|---------------------------|-------------------------|--------------------------|
| a) 3427 | b) 0.00456 | c) 123,453 |
| d) 172 | e) 0.000984 | f) 0.502 |
| g) 3100.0×10^2 | h) 0.0114×10^4 | i) 107.2 |
| j) 0.0000455 | k) 2205.2 | l) 30.0×10^{-2} |
| m) 0.982×10^{-3} | n) 0.0473 | o) 650.502 |
| p) 3.03×10^{-1} | q) 20.4×10^5 | r) 1.29 |
| s) 0.00565 | t) 1362205.2 | u) 450.0×10^3 |
| v) 1000×10^{-3} | w) $546,000 \pm 10$ | x) $546,000 \pm 1000$ |

3. Convert each into decimal form.

- | | | | |
|-----------------------|---------------------------|------------------------|---------------------|
| 1.56×10^4 | 0.56×10^{-2} | 3.69×10^{-2} | 736.9×10^5 |
| 0.00259×10^5 | 0.000459×10^{-1} | 13.69×10^{-2} | 6.9×10^4 |
| 0.00259×10^3 | 0.0209×10^{-3} | | |

4. Calculate the following. Give the answer in correct scientific notation.

- | | | | |
|--|-----------------------------------|---|--|
| a) 4.53×10^5
+ 2.2×10^6 | b) 1913.0
- 4.6×10^3 | c) 2.34×10^{-4}
+ 1.92×10^{-23} | d) 2.130×10^3
- 6.6×10^2 |
| e) 9.10×10^3
+ 2.2×10^6 | f) 1113.0
- 14.6×10^2 | g) 6.18×10^{-45}
+ 4.72×10^{-44} | h) 4.25×10^{-3}
- 1.6×10^{-2} |

5. Calculate the following. Give the answer in correct scientific notation.

- a) $3.95 \times 10^2 / 1.5 \times 10^6$ b) $(3.5 \times 10^2)(6.45 \times 10^{10})$ c) $4.44 \times 10^7 / 2.25 \times 10^5$ d) $(4.50 \times 10^{-12})(3.67 \times 10^{-12})$
- e) $1.05 \times 10^{-26} / 4.2 \times 10^{-26}$ f) $(2.5 \times 10^9)(6.45 \times 10^4)$ g) $6.022 \times 10^{23} / 3.011 \times 10^{-26}$ h) $(6.68 \times 10^5)(3.45 \times 10^{-10})$

Unit 08 Practice Problems Chemistry Key Significant Figures

LL Leslie

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