



Unit: Stoichiometry **Answer key** "Mole-Mole Calculations" - WS #1

Directions: For each of the following problems, show your balanced chemical equation and all of your calculations.

1. How many moles of potassium oxide (K_2O) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:



$$1.52 \text{ mol } K \times \frac{2 \text{ mol } K_2O}{4 \text{ mol } K} = 0.760 \text{ mol } K_2O$$

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:



$$0.25 \text{ mol } Cu_2S \times \frac{2 \text{ mol } Cu}{1 \text{ mol } Cu_2S} = 0.50 \text{ mol } Cu$$

2. How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:



$$3.75 \text{ mol } Na_2SO_4 \times \frac{2 \text{ mol } Na}{1 \text{ mol } Na_2SO_4} = 7.50 \text{ mol } Na$$

3. How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:



$$1.6 \text{ mol } Fe_2O_3 \times \frac{3 \text{ mol } CaO}{1 \text{ mol } Fe_2O_3} = 4.8 \text{ mol } CaO$$

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: **magnesium chlorate $\rightarrow$ magnesium chloride + oxygen**

Formula equation: **$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$**
(must be balanced!)

$$5.12 \text{ mol } Mg(ClO_3)_2 \times \frac{3 \text{ mol } O_2}{1 \text{ mol } Mg(ClO_3)_2} = 15.4 \text{ mol } O_2$$

5. How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: **chromium + lead (IV) nitrate $\rightarrow$ chromium (III) nitrate + lead**

Formula equation: **$4Cr + 3Pb(NO_3)_4 \rightarrow 4Cr(NO_3)_3 + 3Pb$**
(must be balanced!)

$$0.85 \text{ mol } Pb(NO_3)_4 \times \frac{4 \text{ mol } Cr(NO_3)_3}{3 \text{ mol } Pb(NO_3)_4} = 1.1 \text{ mol } Cr(NO_3)_3$$

Unit 8 Stoichiometry Test Review Answers

Christian Drosten



Unit 8 Stoichiometry Test Review Answers:

Unit 8 Stoichiometry Test Review Answers: Bestsellers in 2023 The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous captivating novels enthralling the hearts of readers worldwide. Lets delve into the realm of top-selling books, exploring the engaging narratives that have charmed audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This heartfelt tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can prevail. Unit 8 Stoichiometry Test Review Answers : Taylor Jenkins Reid's "The Seven Husbands of Evelyn Hugo" This spellbinding historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reid's captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Unit 8 Stoichiometry Test Review Answers : Delia Owens' "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting. These top-selling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is an exceptional and thrilling novel that will keep you speculating until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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Table of Contents Unit 8 Stoichiometry Test Review Answers

1. Understanding the eBook Unit 8 Stoichiometry Test Review Answers
 - The Rise of Digital Reading Unit 8 Stoichiometry Test Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 8 Stoichiometry Test Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 8 Stoichiometry Test Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 8 Stoichiometry Test Review Answers
 - Personalized Recommendations
 - Unit 8 Stoichiometry Test Review Answers User Reviews and Ratings
 - Unit 8 Stoichiometry Test Review Answers and Bestseller Lists
5. Accessing Unit 8 Stoichiometry Test Review Answers Free and Paid eBooks
 - Unit 8 Stoichiometry Test Review Answers Public Domain eBooks
 - Unit 8 Stoichiometry Test Review Answers eBook Subscription Services
 - Unit 8 Stoichiometry Test Review Answers Budget-Friendly Options
6. Navigating Unit 8 Stoichiometry Test Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 8 Stoichiometry Test Review Answers Compatibility with Devices
 - Unit 8 Stoichiometry Test Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 8 Stoichiometry Test Review Answers
 - Highlighting and Note-Taking Unit 8 Stoichiometry Test Review Answers
 - Interactive Elements Unit 8 Stoichiometry Test Review Answers
8. Staying Engaged with Unit 8 Stoichiometry Test Review Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 8 Stoichiometry Test Review Answers
9. Balancing eBooks and Physical Books Unit 8 Stoichiometry Test Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 8 Stoichiometry Test Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 8 Stoichiometry Test Review Answers
- Setting Reading Goals Unit 8 Stoichiometry Test Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 8 Stoichiometry Test Review Answers
- Fact-Checking eBook Content of Unit 8 Stoichiometry Test Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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