

- 1) The value of ΔG° at 141.0 °C for the formation of phosphorous trichloride from its constituent elements. 1) _____



is _____ kJ/mol. At 25.0°C for this reaction, ΔH° is -720.5 kJ/mol, ΔG° is -642.9 kJ/mol, and ΔS° is -263.7 J/K.

- A) -683.3 B) 3.65×10^4 C) -829.7 D) -612.3 E) 1.08×10^5
- 2) How many grams of copper will be plated out by a current of 2.3 A applied for 25 minutes to a 0.50 M solution of copper(II) sulfate? 2) _____
- A) 1.8×10^{-2} B) 2.2 C) 0.019 D) 0.036 E) 1.1
- 3) The standard cell potential (E°) of a voltaic cell constructed using the cell reaction below is 0.76 V: 3) _____



With $P_{\text{H}_2} = 1.0 \text{ atm}$ and $[\text{Zn}^{2+}] = 1.0 \text{ M}$, the cell potential is 0.66 V. The concentration of H^+ in the cathode compartment is _____ M.

- A) 1.4×10^{-1}
 B) 4.9×10^1
 C) 4.2×10^{-4}
 D) 1.0×10^{-12}
 E) 2.0×10^{-2}
- 4) ΔS is negative for the reaction _____ 4) _____
- A) $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{g})$
 B) $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
 C) $\text{NH}_4\text{Cl}(\text{s}) \rightarrow \text{NH}_3(\text{g}) + \text{HCl}(\text{g})$
 D) $2\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g})$
 E) $\text{PbCl}_2(\text{s}) \rightarrow \text{Pb}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq})$

- 5) The standard emf for the cell using the overall cell reaction below is +2.20 V: 5) _____



The emf generated by the cell when $[\text{AlI}_3] = 4.5 \times 10^{-3} \text{ M}$ and $[\text{I}^-] = 0.15 \text{ M}$ is _____ V.

- A) 2.23 B) 2.32 C) 2.20 D) 2.10 E) 2.39
- 6) Which substance is the reducing agent in the reaction below? 6) _____



- A) H_2SO_4 B) PbO_2 C) Pb D) H_2O E) PbSO_4

Unit 9 Thermochemistry Test

Ted Lister, Janet Renshaw



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Table of Contents Unit 9 Thermochemistry Test

1. Understanding the eBook Unit 9 Thermochemistry Test
 - The Rise of Digital Reading Unit 9 Thermochemistry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 9 Thermochemistry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 9 Thermochemistry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 9 Thermochemistry Test
 - Personalized Recommendations
 - Unit 9 Thermochemistry Test User Reviews and Ratings
 - Unit 9 Thermochemistry Test and Bestseller Lists
5. Accessing Unit 9 Thermochemistry Test Free and Paid eBooks
 - Unit 9 Thermochemistry Test Public Domain eBooks
 - Unit 9 Thermochemistry Test eBook Subscription Services
 - Unit 9 Thermochemistry Test Budget-Friendly Options

6. Navigating Unit 9 Thermochemistry Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 9 Thermochemistry Test Compatibility with Devices
 - Unit 9 Thermochemistry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 9 Thermochemistry Test
 - Highlighting and Note-Taking Unit 9 Thermochemistry Test
 - Interactive Elements Unit 9 Thermochemistry Test
8. Staying Engaged with Unit 9 Thermochemistry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 9 Thermochemistry Test
9. Balancing eBooks and Physical Books Unit 9 Thermochemistry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 9 Thermochemistry Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 9 Thermochemistry Test
 - Setting Reading Goals Unit 9 Thermochemistry Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 9 Thermochemistry Test
 - Fact-Checking eBook Content of Unit 9 Thermochemistry Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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