

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{2.8\text{g}/4.0\text{g}}{1} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol} \\ 250\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} = 6.857\text{mol HCl} \quad \text{2mol in rxn} = 3.43\text{mol}$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{142.0\text{g}/172.02\text{g}}{1} \right] \times 100 = 82.55\%$$

Unit Vii Stoichiometry Hebden Answer

David Callaghan

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Table of Contents Unit Vii Stoichiometry Hebden Answer

1. Understanding the eBook Unit Vii Stoichiometry Hebden Answer
 - The Rise of Digital Reading Unit Vii Stoichiometry Hebden Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit Vii Stoichiometry Hebden Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit Vii Stoichiometry Hebden Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit Vii Stoichiometry Hebden Answer
 - Personalized Recommendations

- Unit Vii Stoichiometry Hebden Answer User Reviews and Ratings
- Unit Vii Stoichiometry Hebden Answer and Bestseller Lists
- 5. Accessing Unit Vii Stoichiometry Hebden Answer Free and Paid eBooks
 - Unit Vii Stoichiometry Hebden Answer Public Domain eBooks
 - Unit Vii Stoichiometry Hebden Answer eBook Subscription Services
 - Unit Vii Stoichiometry Hebden Answer Budget-Friendly Options
- 6. Navigating Unit Vii Stoichiometry Hebden Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit Vii Stoichiometry Hebden Answer Compatibility with Devices
 - Unit Vii Stoichiometry Hebden Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit Vii Stoichiometry Hebden Answer
 - Highlighting and Note-Taking Unit Vii Stoichiometry Hebden Answer
 - Interactive Elements Unit Vii Stoichiometry Hebden Answer
- 8. Staying Engaged with Unit Vii Stoichiometry Hebden Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit Vii Stoichiometry Hebden Answer
- 9. Balancing eBooks and Physical Books Unit Vii Stoichiometry Hebden Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit Vii Stoichiometry Hebden Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit Vii Stoichiometry Hebden Answer
 - Setting Reading Goals Unit Vii Stoichiometry Hebden Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit Vii Stoichiometry Hebden Answer
 - Fact-Checking eBook Content of Unit Vii Stoichiometry Hebden Answer

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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