



Name: \_\_\_\_\_ Date: \_\_\_\_\_

## Student Exploration: Colligative Properties

**Background:** Colligative properties are properties depending on concentration. Colligative properties include boiling point, freezing point, osmotic pressure, and vapor pressure. Colligative properties depend on solute particles.

**Prior Knowledge:** Questions (Do this a few days before the lesson.)

- What is the difference between a homogeneous and a heterogeneous mixture? Why? Give an example.
- Boiling point is the temperature at which a liquid turns into a gas. What does it depend on?

- Antifreeze is added to water to prevent the water from freezing. Antifreeze's freezing point is lower than water's. How does this work? What are the colligative properties?

Antifreeze lowers the freezing point of water to get vehicles to start in the winter. It also keeps your engine warmer, and the antifreeze also helps the water in the radiator the summer when it's really hot. Antifreeze is a solute in the engine coolant solution.

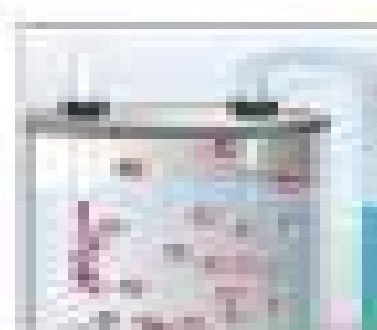
**Warm-up Questions**

Antifreeze is an alcohol substance. For every mole of antifreeze, there are 2 moles of water. Three colligative properties, boiling point, freezing point, and osmotic pressure, are explained in this Colligative Properties Gizmo.

Glenn's chemistry partner is curious about the boiling, or BP, of:

- Water: the chemical formula for water molecules is  $H_2O$ .

Glenn knows that at sea level water boils at 100°C. He is curious about boiling point changes.



# Student Exploration Colligative Properties Gizmo Answers

**Karin Nielsen-Saines**



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