

Guided Practice - Acids & Bases**Properties of Acids and Bases**

1. Match the following terms with their descriptions:

- | | | |
|------------------------|-------|--|
| a. Arrhenius Acid | _____ | the substance formed when a proton is added to _____ |
| b. Arrhenius Base | _____ | _____ a base |
| c. Conjugate Acid | _____ | the remaining substance when a proton is lost from an acid |
| d. Conjugate Base | _____ | a substance that produces OH^- in solution |
| e. Bronsted-Lowry Acid | _____ | a substance producing H^+ in solution |
| f. Bronsted-Lowry Base | _____ | a proton acceptor |
| | | a proton donor |

2. Which of the following represent conjugate acid-base pairs? Label each species as a conjugate acid or a conjugate base. For those pairs that are H_2O conjugates, write the correct conjugate acid or base for each species in the pair, labeling each species correctly.

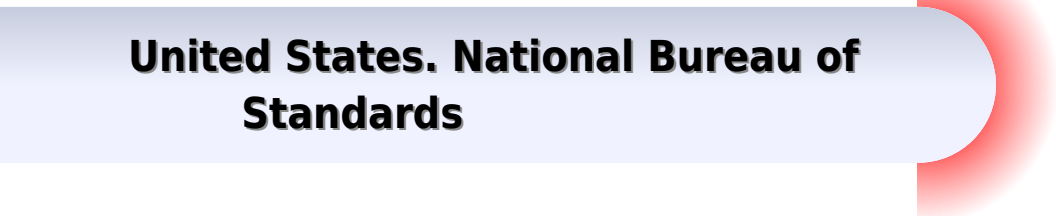
- HSO_4^- , SO_4^{2-}
- HBr , Br^-
- H_3PO_4 , PO_4^{3-}
- H_2O , HO^-

3. Complete the following table regarding acid strength.

Property	Strong Acid	Weak Acid
Definition		
Describe Dissociation		
Species present in solution		
Relative strength of conjugate base compared to water		
Equation showing acid's reaction to water		

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Acid-base Behavior in Aprotic Organic Solvents Marion Maclean Davis, 1968 A unified picture of acid base behavior in aprotic organic solvents is presented based on an extensive survey of the literature and experimental results of the author and associates Evidence given to support this picture includes data pertaining to colligative properties of acids bases and salts and also conductance dielectric constants distribution between immiscible solvents and spectral absorption in the infrared visible and ultraviolet The acids upon which attention is centered are proton donor compounds that are measurably ionized in water such as aliphatic and aromatic carboxylic acids substituted phenols and mineral acids The bases of principal interest are likewise compounds capable of forming ions in water for example aliphatic and aromatic amines and derivatives of guanidine or pyridine The

solvents emphasized are hydrocarbons and halohydrocarbons but data for dipolar aprotic solvents for example acetone acetonitrile and nitrobenzene are included Contrasts in acid base behavior and in acidity and basicity scales in aprotic and water like solvents are discussed The role of hydrogen bonding in aprotic solvents is discussed at length Important types of hydrogen bonded structures include chelate rings self associated acids bases and salts hydrogen bonded ion pairs and homo and heteroconjugate cations and anions Examples are given in which hydrogen bonding of these types affects such properties as the absorption spectrum of a salt the catalytic effect of an acid and the accurate location of a titration endpoint Author

Acid-Base Interactions Kashmiri Lal Mittal, Herbert R. Anderson, 1991-12 This book chronicles the proceedings of the Symposium on Acid Base Interactions Relevance to Adhesion Science and Technology held on the occasion of the 75th birthday of Professor Frederick M Fowkes as a part of the 64th Colloid and Surface Science Symposium held at Lehigh University June 18 20 1990 The book contains 22 papers which are divided into three sections Topics covered include Acid base concepts historical account current status and prospects for the future quantum mechanical approach to understanding acid base interactions at metal polymer interfaces assessment of acidbase interactions at solid liquid interfaces quantitative characterization of the acid base properties of solvents polymers and inorganic surfaces overview by Professor Fowkes acid base characteristics of a variety of solid materials clay minerals carbon fibers glass fibers silicas metals polymers acid base interactions in wetting applications of acid base interactions in a variety of situations e g in the adhesion of polymers to metallic and ceramic substrates mechanical properties of wood properties of filled polymers and behavior of fiber reinforced polymer composites *Hard and Soft Acids and Bases* Ralph G. Pearson, 1973 **A Complete Crash Course in AIEEE** 2011 K.K. Arora, Dinesh Khattar, Ravi Raj Dudeja, **Advanced Organic Chemistry: Reactions and Mechanisms** Singh, Maya Shankar, 2004 Advanced Organic Chemistry Reactions and Mechanisms covers the four types of reactions substitution addition elimination and rearrangement the three types of reagents nucleophiles electrophiles and radicals and the two effects electroni **NBS Monograph** United States. National Bureau of Standards, 1959 **Handbook of Hydrothermal Technology** K. Byrappa, Masahiro Yoshimura, 2012-12-31 Quartz zeolites gemstones perovskite type oxides ferrite carbon allotropes complex coordinated compounds and many more all products now being produced using hydrothermal technology Handbook of Hydrothermal Technology brings together the latest techniques in this rapidly advancing field in one exceptionally useful long needed volume The handbook provides a single source for understanding how aqueous solvents or mineralizers work under temperature and pressure to dissolve and recrystallize normally insoluble materials and decompose or recycle any waste material The result as the authors show in the book is technologically the most efficient method in crystal growth materials processing and waste treatment The book gives scientists and technologists an overview of the entire subject including Evolution of the technology from geology to widespread industrial use Descriptions of equipment used in the process and how it works Problems involved with the growth of crystals processing of technological materials

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