

KEY

19. Consider the *rate* of the following reaction:

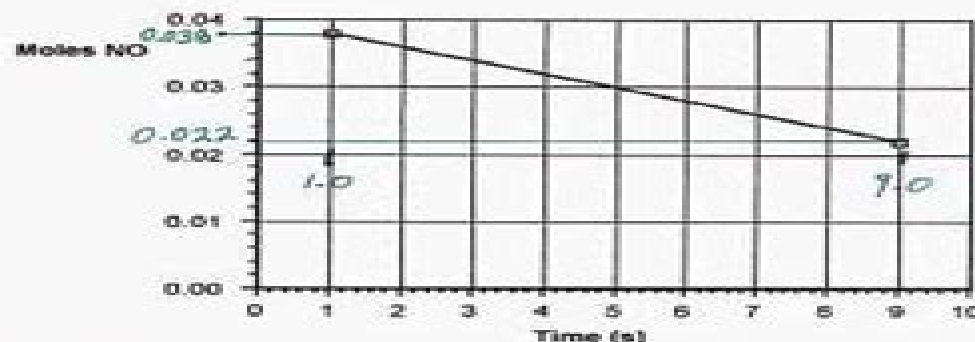


- a) Is it dependent on *temperature*? yes. Explain your answer.
Mostly all reactions are temperature dependent.
- b) Is it dependent on *pressure*? no. Explain your answer.
No gaseous reactants
- c) Is it dependent on *surface area*? yes. Explain your answer.
Heterogeneous reactions. (s) and (aq) reactants

20. Consider the following reaction:



Data collected for the above reaction was used to construct the following graph:



From this graph, determine the *rate of reaction* in moles of NO consumed per second.

$$\text{slope} = - \frac{(0.038 - 0.022) \text{ mol}}{(9.0 - 1.0) \text{ s}} = \frac{0.016 \text{ mol}}{8 \text{ s}} = -0.002 \text{ mol NO/s}$$

$$\text{so rate} = 0.002 \text{ mol NO/s}$$

Answer 0.002 mol NO/s

Topic 8 Kinetics And Equilibrium Answer Key

Michael Hoelscher



Topic 8 Kinetics And Equilibrium Answer Key:

The Enigmatic Realm of **Topic 8 Kinetics And Equilibrium Answer Key**: Unleashing the Language is Inner Magic

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