

Answer to Mole Problems**Mole Worksheet KEY**

Name _____

Period _____

Molar Relationships

Directions: Answer the following questions **on a separate sheet of paper**. Set-up all problems using the factor-label method of dimensional analysis and show all your work and units.

1. What is the mass of 7.50 moles of sulfur dioxide (SO_2)? **480g**

$$\frac{7.5\text{mol}}{1} \times \frac{64\text{g}}{1\text{mol}} = 480\text{g}$$

2. How many moles are there in 250.0 grams of sodium phosphate (Na_3PO_4)? **1.52mol**

$$\frac{250\text{g}}{1} \times \frac{1\text{mol}}{164\text{g}} = 1.52\text{mol}$$

3. How many grams of potassium sulfate (K_2SO_4) are there in 25.3 moles? **4402g**

$$\frac{25.3\text{mol}}{1} \times \frac{176\text{g}}{1\text{mol}} = 4402\text{g}$$

4. What is the volume of 0.38 moles of any gas at STP? **8.5L**

$$\frac{0.38\text{mol}}{1} \times \frac{22.4\text{L}}{1\text{mol}} = 8.5\text{L}$$

7. What is the mass of 51 liters of oxygen gas? **73g**

$$\frac{51\text{L}}{1} \times \frac{1\text{mol}}{22.4\text{L}} \times \frac{32\text{g}}{1\text{mol}} = 73\text{g}$$

1. Calculate the number of moles in 32.2-L of NH_3

$$\frac{32.3\text{L}}{1} \times \frac{1\text{mol}}{22.4\text{L}} = 1.44\text{mol}$$

2. Calculate the number of grams in 3.25-mol of AgNO_3

Using Moles With Formulas Packet Answers

Michael Hoelscher



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Table of Contents Using Moles With Formulas Packet Answers

1. Understanding the eBook Using Moles With Formulas Packet Answers
 - The Rise of Digital Reading Using Moles With Formulas Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Using Moles With Formulas Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Using Moles With Formulas Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Using Moles With Formulas Packet Answers
 - Personalized Recommendations
 - Using Moles With Formulas Packet Answers User Reviews and Ratings
 - Using Moles With Formulas Packet Answers and Bestseller Lists
5. Accessing Using Moles With Formulas Packet Answers Free and Paid eBooks
 - Using Moles With Formulas Packet Answers Public Domain eBooks
 - Using Moles With Formulas Packet Answers eBook Subscription Services
 - Using Moles With Formulas Packet Answers Budget-Friendly Options
6. Navigating Using Moles With Formulas Packet Answers eBook Formats

- ePub, PDF, MOBI, and More
- Using Moles With Formulas Packet Answers Compatibility with Devices
- Using Moles With Formulas Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Using Moles With Formulas Packet Answers
 - Highlighting and Note-Taking Using Moles With Formulas Packet Answers
 - Interactive Elements Using Moles With Formulas Packet Answers
- 8. Staying Engaged with Using Moles With Formulas Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Using Moles With Formulas Packet Answers
- 9. Balancing eBooks and Physical Books Using Moles With Formulas Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Using Moles With Formulas Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Using Moles With Formulas Packet Answers
 - Setting Reading Goals Using Moles With Formulas Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Using Moles With Formulas Packet Answers
 - Fact-Checking eBook Content of Using Moles With Formulas Packet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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