

Annotated periodic table of the elements

1		2		Group name*																13		14		15		16		17		18	
1 H 1.01 Hydrogen				1 Hydrogen & the alkali metals																14 Carbon group		15 Nitrogen group		16 Chalcogens		17 Halogens		18 Noble gases			
3 Li 6.94 Lithium		4 Be 9.01 Beryllium		2 Alkaline earth metals																15 Phosphorus group		16 Sulfur group		17 Halogens		18 Noble gases					
11 Na 22.99 Sodium		12 Mg 24.31 Magnesium		13 Coinage metals																15 Phosphorus group		16 Sulfur group		17 Halogens		18 Noble gases					
19 K 39.10 Potassium		20 Ca 40.08 Calcium		14 Transition metals																15 Phosphorus group		16 Sulfur group		17 Halogens		18 Noble gases					
37 Rb 85.47 Rubidium		38 Sr 87.62 Strontium		15 Transition metals																15 Phosphorus group		16 Sulfur group		17 Halogens		18 Noble gases					
55 Cs 132.91 Cesium		56 Ba 137.33 Barium		16 Transition metals																15 Phosphorus group		16 Sulfur group		17 Halogens		18 Noble gases					
87 Fr 223 Francium		88 Ra 226 Radium		17 Transition metals																15 Phosphorus group		16 Sulfur group		17 Halogens		18 Noble gases					
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- Group names*
- Hydrogen & the alkali metals
 - Alkaline earth metals
 - Coinage metals
 - Volatile metals
 - Isotopes
 - Chalcogens
 - Chalcogens
 - Halogens
 - Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
(Allred-Pingree)

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

39 La 138.91 Lanthanum	40 Ce 140.12 Cerium	41 Pr 140.91 Praseodymium	42 Nd 144.24 Neodymium	43 Pm 144.91 Promethium	44 Sm 150.36 Samarium	45 Eu 151.96 Europium	46 Gd 157.25 Gadolinium	47 Tb 158.93 Terbium	48 Dy 162.50 Dysprosium	49 Ho 164.93 Holmium	50 Er 167.26 Erbium	51 Tm 168.93 Thulium	52 Yb 173.05 Ytterbium	
53 Ag 107.87 Silver	54 Cd 112.41 Cadmium	55 In 114.82 Indium	56 Sn 118.71 Tin	57 Pb 207.2 Lead	58 Bi 208.98 Bismuth	59 Po 209 Polonium	60 At 210 Astatine	61 Rn 222 Radon	62 Ac 227 Actinium	63 Th 232.04 Thorium	64 Pa 231.04 Protactinium	65 U 238.03 Uranium	66 Np 237.05 Neptunium	67 Pu 244.06 Plutonium

Radioactive

Approximate
likely to be either
a metal or a
post-transition
metal

Active metals

- Pre-transition**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphoteric.
- Actinides**
Soft, dense and reactive metals. Those from Th onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

- Transition metals**
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphoteric, and anionic species such as zincates, stannates, and bismuthates.

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- Group 1**
The alkali metals Li, Na, K, Rb, Cs, and Fr are soft, silvery-white metals. They are highly reactive, especially with water, and are stored under oil. They are the most reactive of the elements.

- Covalent nonmetals**
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically covalent nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Word Wise About The Periodic Table And Answers

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