

CHEMISTRY 11 – WORD EQUATIONS AND TYPES OF CHEMICAL REACTIONS WORKSHEET

- 1) Translate the following word equations into chemical symbols and balance the resulting equations. Do not include the phases.

(a) Potassium + water $\rightarrow$ potassium hydroxide + hydrogen



(b) Nitrogen + hydrogen $\rightarrow$ ammonia



- 2) In each of the following write a balanced chemical equation, including the phases.

(a) Liquid phosphoric acid reacts with aqueous barium hydroxide to give water and a precipitate of barium phosphate



(b) Solid aluminium oxide and aqueous sulphuric acid produce water and aqueous aluminium sulphate



(c) Nitrogen trifluoride gas and hydrogen gas react to form nitrogen gas and gaseous hydrogen fluoride



(d) Powdered sodium carbonate and aqueous hydrobromic acid react to form carbon dioxide gas, aqueous sodium bromide and water



- 3) Classify the following reactions.

Reaction	Type
$2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$	Combustion
$\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$	Single Replacement
$4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$	Synthesis
$\text{Sn}(\text{OH})_4 + 4\text{HBr} \rightarrow 4\text{H}_2\text{O} + \text{SnBr}_4$	Neutralization
$2\text{Fe}(\text{NO}_3)_3 + 3\text{MgS} \rightarrow \text{Fe}_2\text{S}_3 + 3\text{Mg}(\text{NO}_3)_2$	Double Replacement
$2\text{N}_2\text{O} \rightarrow 2\text{N}_2 + \text{O}_2$	Decomposition

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Table of Contents Word Equations Chemistry Answers If8767

1. Understanding the eBook Word Equations Chemistry Answers If8767
 - The Rise of Digital Reading Word Equations Chemistry Answers If8767
 - Advantages of eBooks Over Traditional Books
2. Identifying Word Equations Chemistry Answers If8767
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Word Equations Chemistry Answers If8767
 - User-Friendly Interface
4. Exploring eBook Recommendations from Word Equations Chemistry Answers If8767
 - Personalized Recommendations
 - Word Equations Chemistry Answers If8767 User Reviews and Ratings
 - Word Equations Chemistry Answers If8767 and Bestseller Lists
5. Accessing Word Equations Chemistry Answers If8767 Free and Paid eBooks

- Word Equations Chemistry Answers If8767 Public Domain eBooks
- Word Equations Chemistry Answers If8767 eBook Subscription Services
- Word Equations Chemistry Answers If8767 Budget-Friendly Options
- 6. Navigating Word Equations Chemistry Answers If8767 eBook Formats
 - ePub, PDF, MOBI, and More
 - Word Equations Chemistry Answers If8767 Compatibility with Devices
 - Word Equations Chemistry Answers If8767 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Word Equations Chemistry Answers If8767
 - Highlighting and Note-Taking Word Equations Chemistry Answers If8767
 - Interactive Elements Word Equations Chemistry Answers If8767
- 8. Staying Engaged with Word Equations Chemistry Answers If8767
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Word Equations Chemistry Answers If8767
- 9. Balancing eBooks and Physical Books Word Equations Chemistry Answers If8767
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Word Equations Chemistry Answers If8767
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Word Equations Chemistry Answers If8767
 - Setting Reading Goals Word Equations Chemistry Answers If8767
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Word Equations Chemistry Answers If8767
 - Fact-Checking eBook Content of Word Equations Chemistry Answers If8767
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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