

9 page workbook

Covers how covalent bonds work, naming covalent bonds, and drawing Lewis Dot structures of covalent bonds.

What are covalent bonds?

Covalent bonds are bonds in which one or more pairs of valence electrons are shared between the nuclei of both atoms.

A compound is produced when two or more different elements combine covalently. When atoms bond, they produce molecules. A particle in which a covalent compound exists is the same compound. Molecules have a definite composition and ratio. H₂O and H₂O₂ are different compounds.

Two atoms of the same element can have different properties. For example, diamond and graphite are both made of carbon atoms but have different properties.

Covalent

When two

have a

Naming Covalent Compounds

You need to know number prefixes to name covalent bonds.

One = mono-

Two = di-

Three = tri-

Drawing Covalent Bonds

To draw covalent bonds, we use Lewis Dot structures. For each atom in the molecule, use a different color or symbol to represent the valence electrons participating in the bond.

Example: MgO is a molecule made with one atom of magnesium and one atom of oxygen.

Magnesium has 2 valence electrons, and its Lewis dot structure is:



Oxygen has 6 valence electrons, and its Lewis dot structure is:



The Lewis Dot structure for MgO is:



Example: CO₂ is a molecule made with one atom of carbon and two atoms of oxygen.

Oxygen has 6 valence electrons, and its Lewis dot structure is:



element is 2
is to the
the second
atom of
oxygen.
the second
side.

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