Chapter 16 Redox chemistry and corrosion

QL.

Identify each of the following half equations as involving either oxidation or reduction:

- $n = Nn(s) \rightarrow Nn^*(nq) + e^-$
- b Cl₂(g) + 2e⁻ → 2CΓ (aq)
- $c = S(s) + 2e \rightarrow S^{2}(aq)$
- d $Zn(s) \rightarrow Zn^{2+}(aq) + 2e$

A.L.

- a oxidation
- b reduction
- e reduction
- d oxidation

Q2.

Balance the following half equations and then identify each as representing either an oxidation or a reduction reaction:

- a $Fe(s) \rightarrow Fe^{3s}(aq)$
- b $K(s) \rightarrow K^*(aq)$
- $e \to F_2(g) \to F_2(aq)$
- d $O_2(g) \rightarrow O^2$ (aq)

A.2.

- a Fe(s) → Fe^{3*}(aq) + 3e⁻ oxidation
- b K(s) → K⁺(aq) + e⁻ oxidation
- e $F_2(g) + 2e^- \rightarrow 2F^-(aq)$ reduction
- d $O_2(g) + 4e^- \rightarrow 2O^2$ (aq) reduction

Q3.

Iron reacts with hydrochloric acid according to the ionic equation:

 $Fe(s) + 2H^*(aq) \rightarrow Fe^{2*}(aq) + H_2(g)$

- a What has been oxidised in this reaction? What is the product?
- b Write a half equation for the oxidation reaction.
- Identify the oxidant.
- d What has been reduced in this reaction? What is the product?
- Write a half equation for the reduction reaction.
- Identify the reductant.

A.3.

- a Fe(s) has been oxidised to Fe^{2*}(aq)
- b $Fe(s) \rightarrow Fe^{2s}(aq) + 2e$
- c H*(aq)
- d H*(aq) has been reduced to H₂
- e $2H^*(aq) + 2e \rightarrow H_2(g)$
- f Fe(s)

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