

HALF-LIFE PROBLEMS

Name _____ Block _____

1. An isotope of cesium (cesium-137) has a half-life of 30 years. If 1.0 g of cesium-137 disintegrates over a period of 90 years, how many g of cesium-137 would remain?

$$\frac{90}{30} = 3 \text{ } \frac{1}{2} \text{ k}$$

$$\begin{array}{r} 1.0 \\ \times 0.5 \\ \hline 0.5 \\ \times 0.5 \\ \hline 0.25 \end{array}$$

2. Actinium-226 has a half-life of 29 hours. If 100 mg of actinium-226 disintegrates over a period of 58 hours, how many mg of actinium-226 will remain?

$$\frac{58}{29} = 2 \text{ } \frac{1}{2} \text{ k}$$

$$\begin{array}{r} 100 \\ \times 0.5 \\ \hline 50 \\ \times 0.5 \\ \hline 25 \end{array}$$

3. Sodium-25 was to be used in an experiment, but it took 3.0 minutes to get the sodium from the reactor to the laboratory. If 5.0 mg of sodium-25 was removed from the reactor, how many mg of sodium-25 were placed in the reaction vessel 3.0 minutes later if the half-life of sodium-25 is 60 seconds?

$$\frac{3.0}{1.0} = 3 \text{ } \frac{1}{2} \text{ k}$$

$$\begin{array}{r} 5 \text{ mg} \\ \times 0.5 \\ \hline 2.5 \\ \times 0.5 \\ \hline 1.25 \\ \times 0.5 \\ \hline 0.625 \end{array}$$

4. The half-life of isotope X is 2.0 years. How many years would it take for a 4.0 mg sample of X to decay and have only 0.50 mg of it remain?

$$3 \times 2.0 \text{ yr} = 6 \text{ yr}$$

$$\begin{array}{r} 4.0 \\ \times 0.5 \\ \hline 2.0 \\ \times 0.5 \\ \hline 1.0 \\ \times 0.5 \\ \hline 0.5 \end{array} \quad 3 \text{ } \frac{1}{2} \text{ k}$$

5. Selenium-83 has a half-life of 25.0 minutes. How many minutes would it take for a 10.0 mg sample to decay and have only 1.25 mg of it remain?

$$3 \times 25 = 75 \text{ min}$$

$$\begin{array}{r} 10 \\ \times 0.5 \\ \hline 5 \\ \times 0.5 \\ \hline 2.5 \\ \times 0.5 \\ \hline 1.25 \end{array} \quad 3 \text{ } \frac{1}{2} \text{ k}$$

6. The half-life of Po-218 is three minutes. How much of a 2.0 gram sample remains after 15 minutes? Suppose you wanted to buy some of this isotope, and it required half an hour for it reach you. How much should you order if you need to use 0.10 gram of this material?

$$\frac{15}{3} = 5 \text{ } \frac{1}{2} \text{ k}$$

$$\begin{array}{r} 2.0 \\ \times 0.5 \\ \hline 1.0 \\ \times 0.5 \\ \hline 0.5 \\ \times 0.5 \\ \hline 0.25 \\ \times 0.5 \\ \hline 0.125 \\ \times 0.5 \\ \hline 0.0625 \end{array}$$

$$30 \text{ min} : 10 \text{ h.l.}$$

$$1 \times 2 \times 2 \times 2 \times 2 = 2 \times 2 \times 2 \times 2 \times 2 = 102.4 \text{ g}$$

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Table of Contents Worksheet 12 Math Skills Half Life Problem And Answers

1. Understanding the eBook Worksheet 12 Math Skills Half Life Problem And Answers
 - The Rise of Digital Reading Worksheet 12 Math Skills Half Life Problem And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Worksheet 12 Math Skills Half Life Problem And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Worksheet 12 Math Skills Half Life Problem And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Worksheet 12 Math Skills Half Life Problem And Answers
 - Personalized Recommendations
 - Worksheet 12 Math Skills Half Life Problem And Answers User Reviews and Ratings
 - Worksheet 12 Math Skills Half Life Problem And Answers and Bestseller Lists

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 - Worksheet 12 Math Skills Half Life Problem And Answers eBook Subscription Services
 - Worksheet 12 Math Skills Half Life Problem And Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Worksheet 12 Math Skills Half Life Problem And Answers Compatibility with Devices
 - Worksheet 12 Math Skills Half Life Problem And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Worksheet 12 Math Skills Half Life Problem And Answers
 - Highlighting and Note-Taking Worksheet 12 Math Skills Half Life Problem And Answers
 - Interactive Elements Worksheet 12 Math Skills Half Life Problem And Answers
8. Staying Engaged with Worksheet 12 Math Skills Half Life Problem And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Worksheet 12 Math Skills Half Life Problem And Answers
9. Balancing eBooks and Physical Books Worksheet 12 Math Skills Half Life Problem And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Worksheet 12 Math Skills Half Life Problem And Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Worksheet 12 Math Skills Half Life Problem And Answers
 - Setting Reading Goals Worksheet 12 Math Skills Half Life Problem And Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Worksheet 12 Math Skills Half Life Problem And Answers
 - Fact-Checking eBook Content of Worksheet 12 Math Skills Half Life Problem And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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