Chemical Bonds

Section 1 Stability in Bonding

A. Some elements combine chemically and no longer have the same properties they did before forming a compound.

Underlined words and phrases are to be filled in by students on the Note-taking Worksheet.

- B. A(n) chemical formula is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
- C. Atoms form compounds when the compound is more stable than the separate atoms.
 - Noble gases are more chemically stable than other elements because they have a complete outer energy level.
 - 2. Elements that do not have full outer energy levels are more stable in compounds.
 - 3. Atoms can lose, gain, or share electrons to get a stable outer energy level.
 - A(n) chemical bond is the force that holds atoms together in a compound.

DISCUSSION QUESTION:

Why do elements form compounds? To become more chemically stable by cetting a complete outer energy level

Types of Bonds Section 2

- A. A(n) ion is a charged particle because it has more or fewer electrons than protons.
 - When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge.
 - When an atom gains an electron, it becomes a negatively charged ion.
- B. An ionic compound is held together by the ionic bond—the force of attraction between opposite charges of the ions.
 - The result of this bond is a(n) neutral compound.
 - The sum of the charges on the ions is zero.
- C. Molecules are neutral particles formed as a result of sharing electrons.
 - A covalent bond is the force of attraction between atoms sharing electrons.
 - 2. Atoms can form double or triple bonds depending on whether they share two or three pairs of electrons.

Stability In Bonding Section 1 Study Guide

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