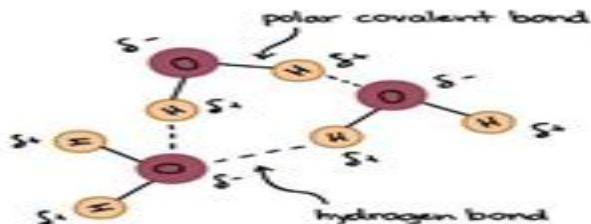


# BIOCHEMISTRY BASICS

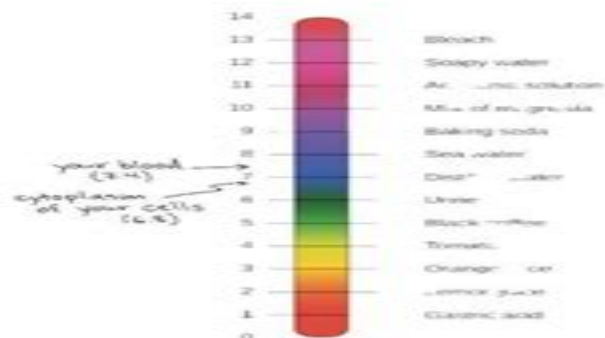
## WATER



- The properties of water relate to its shape, polar nature and its ability to form **hydrogen bonds** with itself and with other molecules.
- Water has a **high specific heat capacity** and a **high heat of vaporisation**.
- Water molecules can ionise into  $H^+$  and  $OH^-$ . **Ionisation** of water is important for living systems because of the reactive nature of  $H^+$ .

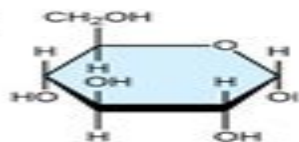
## PH SCALE

- The pH scale is used to express degrees of acidity.
- An **acid** is a substance that causes the concentration of  $H^+$  in a solution to rise above its equilibrium value.
- A **base** is a substance that causes a decrease in the relative numbers of  $H^+$  and an increase in the relative numbers of  $OH^-$  in a solution.
- Buffers** are chemical systems that maintain a **relatively constant pH** even when substantial amounts of acid or base are added.



## BIOLOGICAL MACROMOLECULES

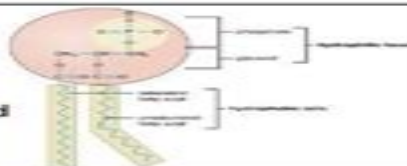
**Carbohydrates** are primarily a combination of carbon and water.



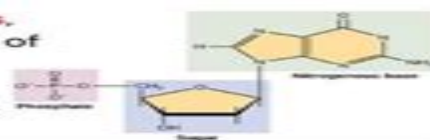
**Proteins** are polymers of amino acids arranged in linear sequence.



**Lipids** are molecules consisting of three fatty acid and glycerol.



**Nucleic acids**, are made out of units called nucleotides (DNA & RNA).



# The Basics Of Biochemistry

**Milin Kurup**



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