

1. In a 15% alcohol solution, the solute is _____ and the solvent is _____.
2. Mixtures characterized by an observable segregation of component substances are called _____ mixtures.
3. The process of crystal decomposition into hydrated ions is called _____.
4. Solutions that are able to dissolve additional solute are said to be _____; solutions that cannot dissolve any more solute at a given temperature are _____; and solutions that contain more solute than can normally be dissolved at a given temperature are _____.
5. Ions that do not participate in a chemical change but are necessary to maintain electrical neutrality in solutions are called _____ ions.
6. Properties that depend only on the number of solute particles are _____ properties.
7. Classify each of the following as a pure substance, a heterogeneous mixture or a homogeneous mixture:
 - a. Smoke
 - b. Diet sprite
 - c. Rubbing alcohol
8. A solution is a _____ mixture.
9. _____ is an example of a solution of gases.
10. Water is an excellent solvent for ionic compounds because each water molecule is _____.
11. Write a balanced dissociation reaction for $\text{Ca}(\text{OH})_2$ dissolved in water.
(remember to use the appropriate subscripts.)
12. What is the difference between a substance like ethanol ($\text{CH}_3\text{CH}_2\text{OH}$) dissolving in water and a substance like sodium chloride (NaCl) dissolving in water?
13. A 0.50 M solution of NaCl is diluted by taking 50.0 mL of it and adding 150.0 mL of water. What is the diluted concentration of the NaCl solution?
14. What mass of CaBr_2 is contained in 75.0 mL of a 0.50 M solution of CaBr_2 ?

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