

25. The number of σ bonds in $N \equiv N$ is
- 1
 - 2
 - 3
 - 4
26. The elements in an ionic compound are held together by
- electrostatic forces of attraction.
 - van der Waals forces
 - the spin of paired electrons.
 - the formation of hybrid orbitals.
 - an electron pair.
27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
- at which oxidation occurs.
 - which attracts cations.
 - at which electrons are supplied to the solution.
 - at which reduction occurs.
28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
- 47
 - 62
 - 93
 - 186
29. In a galvanic (voltaic) cell in which the reaction is $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$ and the ions are at unit concentration (activity), the cell potential is
- $Cd \rightarrow Cd^{2+} + 2e^- \quad 0.4021 \text{ V}$
 $Cu \rightarrow Cu^{2+} + 2e^- \quad -0.344 \text{ V}$
- 0.1383 V
 - 0.4021 V
 - 0.344 V
 - 0.7461 V
 - 0.3677 V
30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
- $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$
 - $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
 - $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$
 - $COCl_2(g) \rightleftharpoons CO(g) + Cl_2(g)$

Standard Potentials	E°
$Mg \rightarrow Mg^{2+} + 2e^-$	2.37V
$Al \rightarrow Al^{3+} + 3e^-$	1.66V
$Zn \rightarrow Zn^{2+} + 2e^-$	0.76V
$Fe \rightarrow Fe^{2+} + 2e^-$	0.44V
$Cu \rightarrow Cu^{2+} + 2e^-$	0.34V
$Ag \rightarrow Ag^+ + e^-$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
- Mg and Ag
 - Mg and Fe
 - Zn and Cu
 - Al and Ag
 - Mg and Al
32. $E = E^\circ - 0.059/n \log Q$ (Nernst equation)
 $[H^+] = 1.0 \text{ M}$ initially, $P_{O_2} = 1.0 \text{ atm}$
- $4e^- + O_2(g) + 4H^+(aq) \rightleftharpoons 2H_2O(l) \quad E^\circ = 1.23 \text{ V}$
 Based on the information above, which statement is correct?
- $n = 1$, since one mole of oxygen is being considered.
 - Addition of base should result in an E value, which is less than 1.23 V.
 - E is independent of the pH of the solution.
 - $Q = \frac{[H_2O]^2}{[O_2][H^+]^4}$
33. The equilibrium constant for the gaseous reaction $C + D \rightleftharpoons E + 2F$ is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
- proceed at equal rates in both directions.
 - proceed more rapidly to form E and F.
 - proceed more rapidly to form C and D.
 - not occur in either direction.
- | Compound | ΔG_f° , kJ/mol |
|-----------|-----------------------------|
| $H_2O(l)$ | -237 |
| $H_2O(g)$ | -229 |
34. At 298 K the equilibrium constant for $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(l)$
- is larger than the K_{eq} for $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(g)$
 - will have a value of 1.0 at equilibrium.
 - cannot be computed since data on O_2 and

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Sir Norman Lockyer



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Table of Contents Study Guide General Chemistry Standardized Exam

1. Understanding the eBook Study Guide General Chemistry Standardized Exam
 - The Rise of Digital Reading Study Guide General Chemistry Standardized Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide General Chemistry Standardized Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide General Chemistry Standardized Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide General Chemistry Standardized Exam
 - Personalized Recommendations
 - Study Guide General Chemistry Standardized Exam User Reviews and Ratings

- Study Guide General Chemistry Standardized Exam and Bestseller Lists
- 5. Accessing Study Guide General Chemistry Standardized Exam Free and Paid eBooks
 - Study Guide General Chemistry Standardized Exam Public Domain eBooks
 - Study Guide General Chemistry Standardized Exam eBook Subscription Services
 - Study Guide General Chemistry Standardized Exam Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Study Guide General Chemistry Standardized Exam Compatibility with Devices
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide General Chemistry Standardized Exam
 - Highlighting and Note-Taking Study Guide General Chemistry Standardized Exam
 - Interactive Elements Study Guide General Chemistry Standardized Exam
- 8. Staying Engaged with Study Guide General Chemistry Standardized Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide General Chemistry Standardized Exam
- 9. Balancing eBooks and Physical Books Study Guide General Chemistry Standardized Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide General Chemistry Standardized Exam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide General Chemistry Standardized Exam
 - Setting Reading Goals Study Guide General Chemistry Standardized Exam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide General Chemistry Standardized Exam
 - Fact-Checking eBook Content of Study Guide General Chemistry Standardized Exam
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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