

< KINETIC THEORY >

Introduction: - (Microscopic properties)

(John Dalton) → Atom → 1803

In 1873 → [Maxwell & Boltzmann]

Ideal Gas → [Kinetic theory of gas]

Atom/molecules constantly motions
They are not in rest.

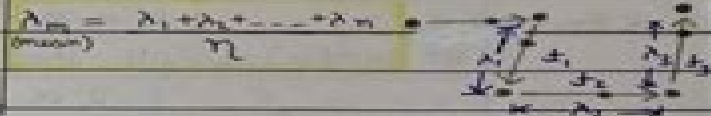
< we study Average properties of gases >

Average Relax time 'T'

- Average time period b/w two successive collisions.

$$T_m = T_1 + T_2 + \dots + T_n$$

(mean) n

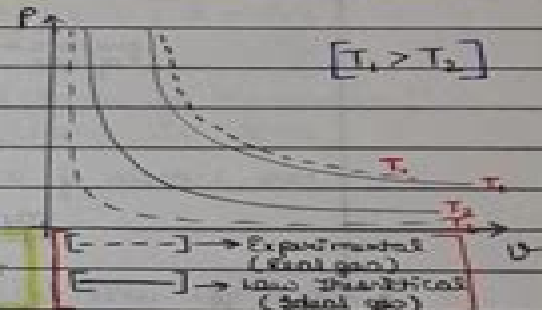


< In 1661 > → [Boyle's Law]: -

- Keeping temperature constant, pressure of a given mass of gas varies inversely with volume.
($P \propto 1/V$)

$$\Rightarrow PV = \text{constant}$$

OR $P_1 V_1 = P_2 V_2$



- At high temperature Real gas and Ideal gas approximately behave same. High temp. per Boyle's law - real, ideal same. per Boyle's law - real, ideal same.

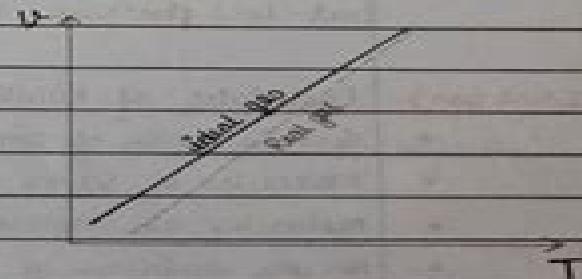
< In 1783 > → [Charles's Law]: -

- For a fixed mass pressure, the vol. of a gas is directly proportional to its absolute temperature 'T'.

$$(V \propto T)$$

$$\Rightarrow \frac{V}{T} = \text{constant}$$

$$\text{OR } \left[\frac{V_1}{T_1} = \frac{V_2}{T_2} \right]$$



- At High Temperature law holds for all gas (Real + Ideal).

Combining both Boyle's and Charles's law

$$(T = \text{constant}) V \propto \frac{1}{P} \quad \cdot \quad V \propto T \quad (P = \text{constant})$$

$$\Rightarrow V \propto \frac{T}{P} \quad \text{OR} \quad \frac{PV}{T} = \text{constant}$$

$$\text{OR } \frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

[Equation of State]

Study Guide For Kinetic Theory

Arshad Iqbal



Study Guide For Kinetic Theory:

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