## **AP Chemistry Final Study Guide**

#### Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- Mass Spectroscopy:
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- Coulomb's Law:
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- · Electromagnetic Radiation:
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- Kinetic Energy of Ionized Electrons:
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so
    it will have more kinetic energy
- · Electron Configuration:
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel: 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>3d<sup>8</sup>4s<sup>2</sup>(notice the d-block)
    - Nickel (with shorthand): [Ar]4s<sup>2</sup>3d<sup>8</sup>
- Law of Definite Proportions:
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- Law of Multiple Proportions:
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- Aufbau Principle:
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# **Study Guide For Second Semester Chemistry Final**

**IM Harris** 

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