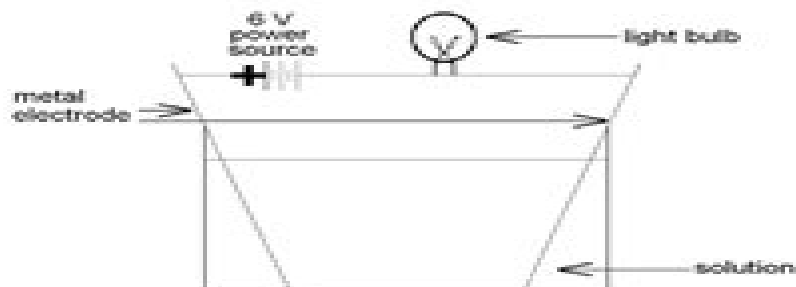


Solutions Unit Review

Multiple-Choice Questions

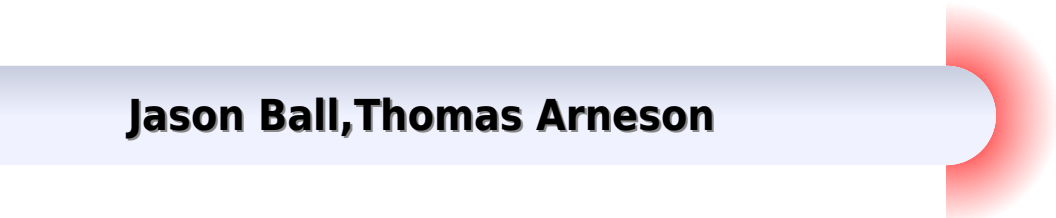
1. All aqueous solutions can be described as
 - a) saturated
 - b) unsaturated
 - c) electrolytes
 - d) homogeneous
2. The apparatus shown in the diagram below is used to determine if an aqueous solution conducts an electric current. Which substance will cause the bulb to light up when added to the water?



- a) $\text{C}_3\text{H}_7\text{OH}(l)$
 - b) $\text{K}_2\text{O}(s)$
 - c) $\text{CH}_3\text{OH}(l)$
 - d) $\text{C}_{12}\text{H}_{22}\text{O}_{11}(s)$
3. A solid is dissolved in water to form an aqueous solution. As dissolving proceeds, the temperature of the solution decreases. What can be concluded from this observation?
 - a) More energy is required to break bonds in the solute than in the solvent.
 - b) More energy is released from bonds forming between solute and solvent than is needed to break bonds in the solute and solvent.
 - c) More energy is needed to break bonds in the solute and solvent than is released from bonds forming between the solute and solvent.
 - d) When the solute dissolves, a net amount of energy is released.
 4. When 1.0×10^{15} formula units of $\text{Ba}(\text{OH})_2$ completely dissolve to form an aqueous solution, the number of ions produced is
 - a) 2.0×10^{15}
 - b) 5.0×10^{15}
 - c) 1.0×10^{15}
 - d) 3.0×10^{15}

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Jason Ball, Thomas Arneson



Solutions Practice Tests:

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