

100. **Continuation**

- (a) Show your answers to (99b) and (99c) to the appropriate number of significant figures. (2)

- (b) Use your answers to (99a) and the relevant equation in page 17 to calculate the number of moles of sodium hydroxide that reacted with 25.0 cm³ of 0.100 mol dm⁻³ sodium hydroxide. (2)

$$\text{moles of NaOH} = \dots\dots\dots \text{mol} \quad (2)$$

- (c) Calculate the concentration of sodium hydroxide in 25.0 cm³.
(Show your working)

$$\text{concentration of NaOH} = \dots\dots\dots \text{mol dm}^{-3} \quad (2)$$

- (d) (i) In this method we measure the percentage of potassium sulfate in the sodium sulfate. (2)

Use your answer to (99b) to show by calculation that the potassium sulfate is in excess.

(2)

- (ii) A student carries out the same method but the concentration of the potassium sulfate solution is too small. (2)

What change to the procedure would the student make to show that the potassium sulfate is in excess?
(Explain your answer.)

.....
.....
..... (2)

(Total: 10)

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