



# CONCENTRATION OF SOLUTIONS

The concentration of a solution is usually measured in moles per cubic decimetre ( $\text{mol dm}^{-3}$ ). This is a measure of the number of moles in one cubic decimetre.

The volume must be in  $\text{dm}^3$  (there are 1000  $\text{cm}^3$  in 1  $\text{dm}^3$ ).  $\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$

$$\text{concentration (mol dm}^{-3}\text{)} = \frac{\text{moles}}{\text{volume (dm}^3\text{)}}$$



1) Calculate the concentration of the following solutions in  $\text{mol dm}^{-3}$ .

a) 0.10 moles of NaCl in 200  $\text{cm}^3$  \_\_\_\_\_

b) 0.20 moles of  $\text{H}_2\text{SO}_4$  in 100  $\text{cm}^3$  \_\_\_\_\_

c) 0.020 moles of NaOH in 25  $\text{cm}^3$  \_\_\_\_\_

2) Calculate the number of moles in the following solutions.

a) 100  $\text{cm}^3$  of 0.20  $\text{mol dm}^{-3}$   $\text{HNO}_3$  \_\_\_\_\_

b) 25  $\text{cm}^3$  of 1.50  $\text{mol dm}^{-3}$  KOH \_\_\_\_\_

c) 50  $\text{cm}^3$  of 0.10  $\text{mol dm}^{-3}$   $\text{H}_2\text{SO}_4$  \_\_\_\_\_

Concentration can also be measured in grams per cubic decimetre ( $\text{g dm}^{-3}$ ). This is a measure of the number of grams in one cubic decimetre. (remember that mass =  $A_r \times \text{moles}$ )

**1  $\text{dm}^3$**

**2 moles of  $\text{H}_2\text{SO}_4$ .**  
**196 g of  $\text{H}_2\text{SO}_4$ .**

Concentration = 2  $\text{mol dm}^{-3}$

MR of  $\text{H}_2\text{SO}_4$  = 98

Concentration = 2  $\times$  98 = 196  $\text{g dm}^{-3}$

A simple conversion is:

$$\text{conc (g dm}^{-3}\text{)} = \text{conc (mol dm}^{-3}\text{)} \times M_r$$

3) Calculate the concentration of the following solutions in  $\text{g dm}^{-3}$ .

a) 0.100  $\text{mol dm}^{-3}$  NaOH \_\_\_\_\_

b) 0.250  $\text{mol dm}^{-3}$   $\text{CH}_3\text{COOH}$  \_\_\_\_\_

c) 1.50  $\text{mol dm}^{-3}$   $\text{HNO}_3$  \_\_\_\_\_

4) 0.20 moles of NaOH is dissolved in 250  $\text{cm}^3$  of water.

a) Calculate the concentration in  $\text{mol dm}^{-3}$  \_\_\_\_\_

b) Calculate the concentration in  $\text{g dm}^{-3}$  \_\_\_\_\_

5) 5.0 g of  $\text{KNO}_3$  is dissolved in 100  $\text{cm}^3$  of water.

a) Calculate the concentration in  $\text{g dm}^{-3}$  \_\_\_\_\_

b) Calculate the concentration in  $\text{mol dm}^{-3}$  \_\_\_\_\_

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**Denise Guinn, Rebecca Brewer, Rachel  
C. Lum**



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