



Troy Bilt Versatiller Manual

Guillaume Favre



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web the coefficient in a balanced equation is an idea the concept of telling the chemist that if the atoms molecules and compounds are balanced there are balanced amounts of the atoms molecules and compounds on the other side in the product

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web the first step to balance the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation after this you can list down the products on the right hand side of the chemical equation

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 $1\text{ N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF}$
 $\text{Cl}_2 + 4\text{H}_2\text{O} \rightarrow 2\text{H}_2 + 4\text{HCl}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + 3\text{Al}_2\text{SO}_4$
 $3\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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 $1\text{ N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{HNO}_3 \rightarrow \text{H}_2\text{O} + 2\text{O}_2 + 3\text{H}_2\text{O}$
 $2\text{H}_2\text{O} + 2\text{H}_2\text{O} \rightarrow 4\text{H}_2 + \text{O}_2$
 $\text{RbH} + 2\text{H}_2\text{O} \rightarrow \text{RbOH} + \text{H}_2$
 $5\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$
 $6\text{N}_2 + 2\text{H}_2 \rightarrow 2\text{NH}_3 + 7\text{N}_2$
 $3\text{Ca} + \text{N}_2\text{O} \rightarrow \text{CaF}_2 + 8\text{CH}_4$
 $2\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{H}_2\text{O} + 9\text{C} + 2\text{H}_5\text{OH}$

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF}$
 $\text{Cl}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $2\text{H}_2\text{O} + \text{Pb} + \text{OH}^- \rightarrow 2\text{HCl} + 2\text{H}_2\text{O} + \text{PbCl}_2$
 $3\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2$
 $\text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2 \rightarrow \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{H}_2\text{O}$

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 $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$
 $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O} + 3\text{HBr}$
 $\text{C}_4\text{H}_{10} + 6\text{O}_3 \rightarrow 1\text{H}_2\text{O} + 2\text{C}_2\text{H}_4\text{O}_2$

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 $38\text{Fe} + 2\text{O}_3 \rightarrow 3\text{CO}_2 + 2\text{Fe}_3\text{CO}_2 + 39\text{H}_2 + 3\text{PO}_4$
 $3\text{BaO} + 3\text{PO}_4 + 2\text{H}_2\text{O} \rightarrow 40\text{Mg} + \text{NH}_4\text{PO}_4 + 4\text{Mg}_2\text{P}_2\text{O}_7 + 2\text{NH}_3 + 2\text{H}_2\text{O}$
 $41\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2\text{SO}_4 + 3\text{H}_2 + 42\text{P}_4 + 10\text{HCl} + 3\text{HPO}_3$
 $43\text{HNO}_2 + 2\text{H}_2\text{O} \rightarrow 3\text{O}_2 + 4\text{HNO}$

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