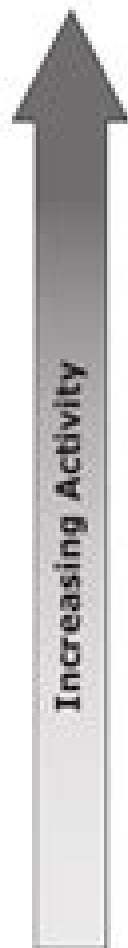


# STAAR CHEMISTRY REFERENCE MATERIALS



POLYATOMIC IONS		SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER		ACTIVITY SERIES
Acetate	$C_2H_3O_2^-$ , $CH_3COO^-$	<b>Soluble.</b> <u>compounds contain</u>	<b>Common exceptions</b>	<b>Metal</b>
Ammonium	$NH_4^+$	$C_2H_3O_2^-$ , $CH_3COO^-$	None	Lithium
Carbonate	$CO_3^{2-}$	$NH_4^+$	None	Potassium
Chlorate	$ClO_3^-$	$NO_3^-$	None	Barium
Chlorite	$ClO_2^-$	$CN^-$	None	Calcium
Chromate	$CrO_4^{2-}$	$ClO^-$	None	Sodium
Cyanide	$CN^-$	$CO_3^{2-}$	None	Magnesium
Dichromate	$Cr_2O_7^{2-}$	$CO_3^{2-}$	None	Aluminum
Hydrogen carbonate	$HCO_3^-$	$CO_3^{2-}$	None	Manganese
Hydroxide	$OH^-$	$Br^-$	Compounds of $Ag^+$ , $Pb^{2+}$ , and $Hg_2^{2+}$	Zinc
Hypochlorite	$ClO^-$	$Cl^-$	Compounds of $Ag^+$ , $Pb^{2+}$ , and $Hg_2^{2+}$	Chromium
Nitrate	$NO_3^-$	$I^-$	Compounds of $Ag^+$ , $Pb^{2+}$ , and $Hg_2^{2+}$	Iron
Nitrite	$NO_2^-$	$SO_4^{2-}$	Compounds of $Sr^{2+}$ , $Ba^{2+}$ , $Pb^{2+}$ , and $Hg_2^{2+}$	Cobalt
Perchlorate	$ClO_4^-$	<b>Insoluble.</b> <u>compounds contain</u>	<b>Common exceptions</b>	Nickel
Permanganate	$MnO_4^-$	$CO_3^{2-}$	Compounds of $NH_4^+$ and the alkali metal cations	Tin
Phosphate	$PO_4^{3-}$	$PO_3^{2-}$	Compounds of $NH_4^+$ and the alkali metal cations	Lead
Sulfate	$SO_4^{2-}$	$CrO_4^{2-}$	Compounds of $NH_4^+$ and the alkali metal cations	(Hydrogen)
Sulfite	$SO_3^{2-}$	$OH^-$	Compounds of $NH_4^+$ and the alkali metal cations, $Ca^{2+}$ , $Sr^{2+}$ , and $Ba^{2+}$	Copper
		$S^{2-}$	Compounds of $NH_4^+$ , the alkali metal cations, $Ca^{2+}$ , $Sr^{2+}$ , and $Ba^{2+}$	Mercury
				Silver
				Platinum
				Gold



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